

Electrochemistry Practice Test

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I. Multiple Choice

1. When a metal undergoes corrosion, it:

- (a) loses electrons
(b) becomes reduced

- (c) acts as an oxidizing agent
(d) decreases in oxidation number

2. At standard conditions, Fe^{+2} reacts spontaneously with

- (a) I_2
(b) Co

- (c) Br^-
(d) Ag^+

3. Which of the following half-reactions is balanced?

- (a) $\text{SO}_4^{-2} + \text{H}_2\text{O} \rightleftharpoons \text{SO}_3^{-2} + 2\text{H}^+ + 2\text{e}^-$
(b) $\text{SO}_4^{-2} + \text{H}_2\text{O} + 2\text{e}^- \rightleftharpoons \text{SO}_3^{-2} + 2\text{H}^+$
(c) $\text{SO}_4^{-2} + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{SO}_3^{-2} + \text{H}_2\text{O}$
(d) $\text{SO}_4^{-2} + 2\text{H}^+ \rightleftharpoons \text{SO}_3^{-2} + \text{H}_2\text{O} + 2\text{e}^-$

4. During a redox reaction, the oxidizing agent:

- (a) reduces other species
(b) gains electrons

- (c) increases in oxidation number
(d) becomes oxidized

5. For a given redox rxn, the oxidation # of tin changed from +2 to +4. As a result, tin:

- (a) lost 2 electrons and was reduced
(b) gained 2 electrons and was reduced
(c) lost 2 electrons and was oxidized
(d) gained 2 electrons and was oxidized

6. In which of the following compounds does carbon have an oxidation number of -2?

- (a) CO
(b) CO_2

- (c) CH_2O
(d) CH_3OH

7. Which of the following equations represents a redox reaction?



8. Consider the following reaction:

Chemistry 12 Electrochemistry Test

**Peter Schwamb, Allyne Litchfield
Merrill**



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