

## Diagnostische toets klas 4 Sk hoofdstuk 4 en 5: Antwoorden

Laatst gewijzigd 18-05-2020

Kijk met een andere kleur na!

OF betekent: ook dit antwoord is goed.

Wat tussen haakjes staat hoeft niet in jouw antwoord te staan.

Eenheid vergeten bij eindantwoord rekenwerk: - 1pt

**Vraag 1:** Voor de pijl mist de  $7\text{H}_2\text{O}(\text{l})$  (1pt)

**Vraag 2:**

1) Onjuist afgerond: 1251,849 moet in één decimaal en het cijfer na de 8 is een 4 dus: 1251,8 (1pt)

2) Eenheden ontbreken: 1251,8 mL (1pt)

**Vraag 3:**

Significantie onjuist:  $2 \times 79,90 = 159,8$  (moet in 4 cijfers (= meetwaarde, de 2 is een telwaarde)) dus één decimaal dus antwoord ook in één decimaal 184,1  $\text{g mol}^{-1}$  (1pt)

**Vraag 4:**

1) De concentratie haken horen niet in een reactie of oplosvergelijking. (1pt)

2) Voor de pijl staat  $\text{H}_2\text{O}$  maar er treedt geen reactie op tussen het magnesiumchloride en het water, er is slechts sprake van oplossen. (1pt)

3) Het gaat niet om het element chloor dus molecuul  $\text{Cl}_2$  maar om gereageerd chloor en er zijn twee chloride-ionen, dus  $2\text{Cl}^-$  (1pt)

4) Er ontbreekt tweemaal (aq) na de pijl bij de ionen. (1pt)

**Vraag 5:**

1) Het natriumion heeft een lading van +1 en niet +2. Dus ook  $\text{Na}_2\text{SO}_4(\text{s})$  en  $2\text{Na}^+(\text{aq})$ . (1pt)

2) Het sulfaation ( $\text{SO}_4^{2-}$ ) is covalent met elkaar verbonden (gedeelde elektronenparen) dus valt niet uitelkaar tijdens het oplossen. (1pt)

**Vraag 6:**

1) In een reactievergelijking komen alleen die stoffen die daadwerkelijk reageren, dus niet de tribune-ionen. (1pt)

2) In een verhoudingsformule wordt geenromeins cijfer (de lading van het metaalion) genoteerd, dat wordt alleen in de naam van het zout gedaan. (1pt)

3) Er zijn twee hydroxide-groepen ( $\text{OH}^-$ ) verbonden met het ijzer(II) ion en dat dient tussen haakjes genoteerd te worden:  $\text{Fe}(\text{OH})_2(\text{s})$  (1pt)

4) Boven de pijl hoeft niet genoteerd te worden 'neerslaan', dat is uit de reactievergelijking wel op te maken dat dat gebeurt. (1pt)

**Vraag 7:**

1) Hier wordt wegen en tellen doorelkaar gehaald, ze moet de molverhouding gebruiken. (1pt)

2) [opgelost deeltje] betekent de molariteit en dat wordt uitgedrukt in mol per liter en niet in gram per milliliter. (1pt)

3) Molair massa magnesiumchloride is 95,211  $\text{g mol}^{-1}$  (Binastabel 98)

$19,75 \text{ g} : 95,211 \text{ g mol}^{-1} = 0,20743 \text{ mol MgCl}_2$  (1pt)

Per liter is dat  $1000/175 \times 0,20743 \text{ mol} = 1,185 \text{ mol MgCl}_2$  (1pt)

Vanwege de molverhouding ( $\text{Mg}^{2+} : \text{Cl}^- = 1 : 2$ )  $1,185 \times 2$  wordt  $[\text{Cl}^-] =$

# Diagnostische Toets Hoofdstuk K

**G Orfield**



## Diagnostische Toets Hoofdstuk K:

**Hacken voor dummies** Oliver-Christopher Rochford, 2004      **Photoshop Klassiekers** S. Kelby, 2006      **Leerling en leraar in samenspraak. De betekenis van onderwijsdialogen voor de zelfstandigheidsontwikkeling van leerlingen in het voortgezet onderwijs** Daan Lockhorst, 2003      **Euclides**, 2007      Hoe geletterd/gecijferd is Vlaanderen? Functionele taal- en rekenvaardigheden van Vlamingen in internationaal perspectief Dirk van Damme, Luc van de Poele, Els Verhasselt, 1997      **Handleiding OBVL** Ad Vermulst, Gert Kroes, Ronald De Meyer, Linda Nguyen, Jan Willem Veerman, 2015-07-17 De Opvoedingsbelastingvragenlijst OBVL maakt deel uit van de set Vragenlijsten Gezin Opvoeding VG O Het gaat om vier vragenlijsten die diverse aspecten van de opvoedingsomgeving in beeld brengen en zelfstandig kunnen worden afgenomen Naast de OBVL maken ook de Vragenlijst Gezinsfunctioneren volgens Ouders VGFO de Verkorte Schaal voor Opvoedersgedrag VSOG en de Vragenlijst Meegemaakte Gebeurtenissen VMG deel uit van de VG O De vragenlijsten zijn ontworpen voor gebruik in instellingen voor jeugdzorg in de brede zin van het woord dus zowel voor instellingen voor jeugd en opvoedhulp als voor de instellingen op het gebied van geestelijke gezondheidszorg justitiele zorg en de zorg voor licht verstandelijk gehandicapten Ook in de jeugdgezondheidszorg het maatschappelijk werk het speciaal onderwijs en de Centra voor Jeugd en Gezin is de VG O set bruikbaar De lijsten beogen een aanvulling en hulpmiddel te zijn voor diagnostiek en evaluatie Alle lijsten worden ingevuld door de primaire opvoeder meestal n van de ouders Deze handleiding gaat over de OBVL deze lijst kent 34 vragen die vijf domeinen van opvoedingsbelasting meten Problemen opvoeder kindrelatie Problemen met opvoeden Depressieve stemmingen Rolbeperking en Gezondheidsklachten Ook is er een totaalscore te berekenen voor de totale door de opvoeder ervaren opvoedingsbelasting De handleiding is een vervolg op de handleiding uit 2012 De lijst zelf en de normen zijn niet gewijzigd alleen de handleiding is verbeterd Tevens is er een hoofdstuk over de korte versie van de OBVL OBVL K toegevoegd      **Slippage Tests** Roelof Doornbos, 1966      Toetsen in het hoger onderwijs Henk van Berkel, Anneke Bax, Desirée Joosten-ten Brinke, 2017-05-11 Dit boek geeft een breed overzicht van toetsen en beoordelingsvormen in het hoger onderwijs Ook beschrijft het verschillende aspecten van toetsen en het toetsproces Het boek geeft handreikingen en praktijkvoorbeelden Daarmee is het een praktische en inspirerende informatiebron voor opleidingen examencommissies toetscommissies en aankomende docenten In deze vierde druk staan nieuwe hoofdstukken over onder meer de kwaliteitspiramide het samenwerken bij toetsontwikkeling het toetsen van creativiteit en over de rechtsbescherming van examinandi De hoofdstukken over digitaal toetsen en ervaringscertificaten EVC zijn herschreven Bijvoorbeeld met extra informatie over flexibel onderwijs In ieder hoofdstuk is veel aandacht voor de rol van de docent Toetsen in het hoger onderwijs kan gebruikt worden om een goed toetsbeleid op te zetten en uit te voeren Daarmee is het geschikt om het certificaat Basiskwalificatie Examinering BKE en het Seniorkwalificatie Examinering SKE te behalen Het boek is toegankelijk en daarom ook goed bruikbaar voor toetsconstructie in het voortgezet onderwijs en middelbaar

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