

5-A-Day Language Review: Week 1

Wednesday

- 1 Circle the **definition** that matches the word used in the sentence below.

re•tire

1. v. To stop working because one has reached a certain age
2. v. To go to bed

I am so exhausted from running the marathon that I am going to retire early tonight.

- 2 Add **commas** where needed to the text below.

Miguel went to the store to buy an onion some lettuce and rice.

- 3 Circle the **conjunction** in the sentence and identify if it is "coordinating" or "subordinating."

Even though it was extremely hot, we went for a 5-mile run.

(C) (S)

- 4 Using your **pronunciation key**, circle the word that rhymes with the bolded word.

bass (bās)

pass

face

mess

rack

- 5 What does the following **adage** or **proverb** mean?

"The early bird gets the worm."

Thursday

- 1 Combine the two sentences with the appropriate **correlative conjunctions**.

both...and • **either...or** • **neither...nor** • **not only...but (also)** • **whether...or**

You can go on the field trip tomorrow.

You can stay in the office tomorrow.

- 2 Circle the **preposition(s)** and cross out the **prepositional phrase(s)**. Underline the **subject** once and the **verb phrase** twice.

Jennifer walked over to her neighbor's house.

- 3 List five words that contain the **Greek** or **Latin root/affix**.

aud - to hear or listen

- 4 Mina agreed to go on the **excursion**, but only because she was getting stir crazy sitting in her hotel room all day. Normally, little trips to see the sights did not interest her. But her teammates convinced her to make the best of her trip and explore the beauty of Mexico.

excursion means: _____

- 5 Choose the **antonym** for: **revolt**

rebel

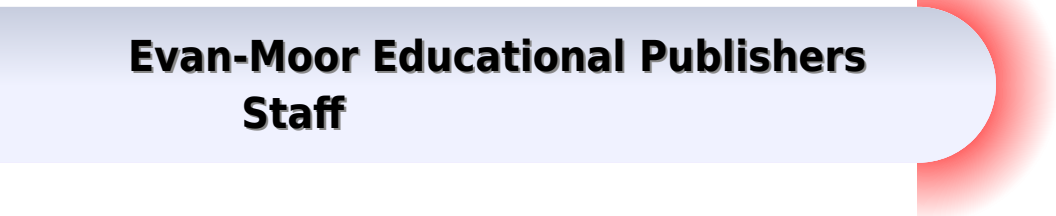
agree

upheaval

resist

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$ $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27.0 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} = 0.593 \text{ mol } \text{Al}_2\text{O}_3$ $0.593 \text{ mol } \text{Al}_2\text{O}_3 \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 60.3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{60.3 \text{ g } \text{Al}_2\text{O}_3}{76.0 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 79.2\%$

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