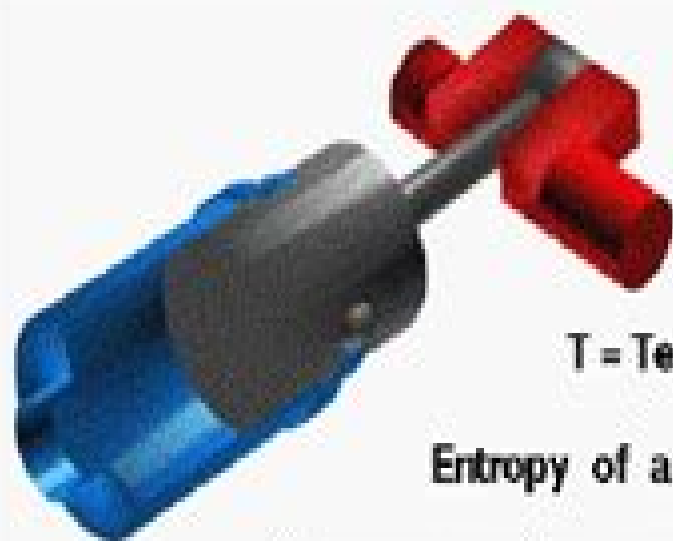




C_p = specific heat at constant pressure

C_v = specific heat at constant volume R = Gas constant

$$C_p - C_v = R \quad \gamma = C_p / C_v$$

Algebra: $R / C_p = 1 - 1 / \gamma = (\gamma - 1) / \gamma$

T = Temperature p = Pressure s = Entropy v = Volume

Entropy of a Gas: $s_2 - s_1 = C_p \ln (T_2 / T_1) - R \ln (p_2 / p_1)$

Compression and expansion are reversible: $s_2 - s_1 = 0$

Then: $C_p \ln (T_2 / T_1) = R \ln (p_2 / p_1)$

Effect of Pressure on Temperature:

$$(T_2 / T_1) = (p_2 / p_1)^{[(\gamma - 1) / \gamma]}$$

Equation of State: $p v = R T$

Substitute: $(p_2 v_2 / p_1 v_1) = (p_2 / p_1)^{[(\gamma - 1) / \gamma]}$

Gather: $v_2 / v_1 = (p_2 / p_1)^{[-1 / \gamma]}$

Effect of Volume on Pressure:

$$p_2 / p_1 = (v_1 / v_2)^\gamma$$

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Yumin Zhang



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