

# THE MOLE AND AVOGADRO'S NUMBER

Name \_\_\_\_\_

One mole of a substance contains Avogadro's Number ( $6.02 \times 10^{23}$ ) of molecules. (atoms)

How many molecules are in the quantities below?

1. 2.0 moles  $x = 2.0 \text{ mol} \times \left( \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{1.2 \times 10^{24} \text{ molecules}}$

2. 1.5 moles  $x = 1.5 \text{ mol} \times \left( \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{23} \text{ molecules}}$

3. 0.75 mole  $x = 0.75 \text{ mol} \times \left( \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{4.5 \times 10^{23} \text{ molecules}}$

4. 15 moles  $x = 15 \text{ mol} \times \left( \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{24} \text{ molecules}}$

5. 0.35 mole  $x = 0.35 \text{ mol} \times \left( \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{2.1 \times 10^{23} \text{ molecules}}$

How many moles are in the number of molecules below?

1.  $6.02 \times 10^{23}$   $x \text{ moles} = 6.02 \times 10^{23} \text{ molecules} \left( \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.0 \text{ mol}}$

2.  $1.204 \times 10^{24}$   $x \text{ moles} = 1.204 \times 10^{24} \text{ molecules} \left( \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.0 \text{ moles}}$

3.  $1.5 \times 10^{20}$   $x \text{ moles} = 1.5 \times 10^{20} \text{ molecules} \left( \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.5 \times 10^{-4} \text{ moles}}$  (0.00024 moles)

4.  $3.4 \times 10^{26}$   $x \text{ moles} = 3.4 \times 10^{26} \text{ molecules} \left( \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{565 \text{ moles}}$

5.  $7.5 \times 10^{19}$   $x \text{ moles} = 7.5 \times 10^{19} \text{ molecules} \left( \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.2 \times 10^{-4} \text{ moles}}$

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**R Sandford**



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