
ACS ORGANIC CHEMISTRY FINAL STUDY GUIDE

1. no pi bonds (unsaturation): 0 degrees
2. 1 pi bond (unsaturation): 1 degree
3. triple bond (unsaturation): 2 degrees
4. formal charge =: # valence e's - (# nonbonding e's + 1/2 # bonding e's)
5. 4 regions of e- density: sp³ hybridization, 109.5 bond angle
6. 3 regions of e- density: sp² hybridization, 120 bond angle
7. 2 regions of e- density: sp hybridization, 180 bond angle
8. aliphatic: Pertaining to any member of one of the two major groups of organic compounds, those having straight or branched chain structures.
9. vinylic: the carbon atoms of a carbon-carbon double bond
10. higher stability =: lower energy
11. higher acidity =: higher conjugate base stability
12. higher acidity =: lower basicity
13. higher pK_a =: stronger base
14. Zaitsev's Rule: The most substituted alkene is formed preferentially
15. Markovnikov: The H atom will always bond to the C with more H
16. diastereomer: Stereoisomers that are not mirror images of one another
17. nucleophilicity: decreases across a period and increases down a group
18. enantiomer: nonsuperimposable mirror images
19. 1680 cm⁻¹ (IR): amide, C=O stretch
20. 1710 cm⁻¹ (IR): ketone, C=O stretch
21. 1710 cm⁻¹ (IR): acid, C=O stretch
22. 1725 cm⁻¹ (IR): aldehyde, C=O stretch
23. 1735 cm⁻¹ (IR): ester, C=O stretch
24. centered 3000 cm⁻¹ (IR): acid, O-H stretch
25. centered 3300 cm⁻¹ (IR): alcohol, O-H stretch
26. 3100-3300 cm⁻¹ (IR): amide/amine, N-H stretch, NH_n where n=# of peaks
27. 2710 & 2810 cm⁻¹ (IR): aldehyde, C-H stretch

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