

Equilibria Test SCH4U Name:

15 K 10 T 11 C 10 A Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (15K)

1. This symbol () indicates that _____

- A) heat must be applied
- B) an incomplete combustion reaction has occurred
- C) a gas is formed by the reaction
- D) the reaction is reversible

2. Which one of the following will change the value of an equilibrium constant?

- A) changing temperature
- B) adding other substances that do not react with any of the species involved in the equilibrium
- C) varying the initial concentrations of reactants
- D) varying the initial concentrations of products

3. The K_p for the reaction below is 1.49×10^2 at 100.0°C : $\text{CO(g)} + \text{Cl}_2\text{(g)} \rightleftharpoons \text{COCl}_2\text{(g)}$

In an equilibrium mixture of the three gases, $P_{\text{CO}} = P_{\text{Cl}_2} = 8.60 \times 10^{-4} \text{ atm}$. The partial pressure of the product, phosgene (COCl_2), is _____ atm.

- A) 1.1×10^2
- B) 2.01×10^{18}
- C) 4.96×10^{-15}
- D) 1.28×10^5
- E) 172×10^{11}

4. Dinitrogen tetroxide partially decomposes according to the following equilibrium:

$\text{N}_2\text{O}_4\text{(g)} \rightleftharpoons 2\text{NO}_2\text{(g)}$ A 1.000-L flask is charged with $3.00 \times 10^{-2} \text{ mol}$ of N_2O_4 . At equilibrium, $2.36 \times 10^{-2} \text{ mol}$ of N_2O_4 remains. K_c for this reaction is _____

- A) 0.723
- B) 0.391
- C) 0.212
- D) 6.94×10^{-3}
- E) 1.92×10^{-3}

5. At 400 K, the equilibrium constant for the reaction $\text{Cl}_{2(g)} + \text{Br}_{2(g)} \rightleftharpoons 2\text{BrCl}_{(g)}$

is $K_p = 7.0$. A closed vessel at 400 K is charged with 1.00 atm of $\text{Br}_2\text{(g)}$, 1.00 atm of $\text{Cl}_2\text{(g)}$, and 2.00 atm of BrCl(g) . Use Q to determine which of the statements below is true

- A) The equilibrium partial pressures of Br_2 , Cl_2 , and BrCl will be the same as the initial values.
- B) The equilibrium partial pressure of Br_2 will be greater than 1.00 atm.
- C) At equilibrium, the total pressure in the vessel will be less than the initial total pressure
- D) The equilibrium partial pressure of BrCl(g) will be greater than 2.00 atm.
- E) The reaction will go to completion since there are equal amounts of Br_2 and Cl_2 .

Chemistry Equilibrium Test Answers

J Dewey



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