Part C: Determination of the molar mass of naphthalene

Average: -3.413 (0

Name of solute: Supposed Mass of solute: 509

Freezing points

Trial 1: 3.3 0

Trial 2: -3.5 0

Trial 3: -3.5 0

1. Calculate the freezing point depression caused by the addition of naphthalene.

2. Calculate the molality of the cyclohexane/naphthalene solution. Use the freezing point depression constant calculated from Part B. — 10.34 = 418.35 x M ☐ 19.26

How many moles of the solute are present? What is the experimental molar mass of the solute?
 Use your equations for molality and freezing point depression to solve for this.

 The molar mass for naphthalene is 128.17 g/mole. Calculate the percent error between your experimental molar mass and the actual molar mass.

Chemistry 1212 Laboratory Manual Experiment 5

Jianjun Gao

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Table of Contents Chemistry 1212 Laboratory Manual Experiment 5

- 1. Understanding the eBook Chemistry 1212 Laboratory Manual Experiment 5
 - The Rise of Digital Reading Chemistry 1212 Laboratory Manual Experiment 5
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry 1212 Laboratory Manual Experiment 5
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 1212 Laboratory Manual Experiment 5
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry 1212 Laboratory Manual Experiment 5
 - Personalized Recommendations

- Chemistry 1212 Laboratory Manual Experiment 5 User Reviews and Ratings
- Chemistry 1212 Laboratory Manual Experiment 5 and Bestseller Lists
- 5. Accessing Chemistry 1212 Laboratory Manual Experiment 5 Free and Paid eBooks
 - Chemistry 1212 Laboratory Manual Experiment 5 Public Domain eBooks
 - Chemistry 1212 Laboratory Manual Experiment 5 eBook Subscription Services
 - Chemistry 1212 Laboratory Manual Experiment 5 Budget-Friendly Options
- 6. Navigating Chemistry 1212 Laboratory Manual Experiment 5 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chemistry 1212 Laboratory Manual Experiment 5 Compatibility with Devices
 - Chemistry 1212 Laboratory Manual Experiment 5 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 1212 Laboratory Manual Experiment 5
 - Highlighting and Note-Taking Chemistry 1212 Laboratory Manual Experiment 5
 - Interactive Elements Chemistry 1212 Laboratory Manual Experiment 5
- 8. Staying Engaged with Chemistry 1212 Laboratory Manual Experiment 5
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - \circ Following Authors and Publishers Chemistry 1212 Laboratory Manual Experiment 5
- 9. Balancing eBooks and Physical Books Chemistry 1212 Laboratory Manual Experiment 5
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 1212 Laboratory Manual Experiment 5
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 1212 Laboratory Manual Experiment 5
 - Setting Reading Goals Chemistry 1212 Laboratory Manual Experiment 5
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 1212 Laboratory Manual Experiment 5
 - Fact-Checking eBook Content of Chemistry 1212 Laboratory Manual Experiment 5

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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