

CHAPTER 9 Chemical Reactions

Section 1 Reactions and Equations

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Practice Problems

pages 284–287

Write skeleton equations for the following word equations.

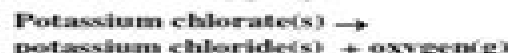
1. Hydrogen and bromine gases react to yield hydrogen bromide.



2. When carbon monoxide and oxygen react, carbon dioxide forms.



3. **Challenge** Write the word equation and the skeleton equation for the following reaction: when heated, solid potassium chlorate yields solid potassium chloride and oxygen gas.



Write chemical equations for each of the following reactions.

4. In water, iron(III) chloride reacts with sodium hydroxide, producing solid iron(III) hydroxide and sodium chloride.



5. Liquid carbon disulfide reacts with oxygen gas, producing carbon dioxide gas and sulfur dioxide gas.



6. **Challenge** A piece of zinc metal is added to a solution of dihydrogen sulfate. This reaction produces a gas and a solution of zinc sulfate.



Section 1 Review

page 288

7. **Explain** why it is important that a chemical equation be balanced.

Because mass is neither created nor destroyed in chemical reactions, the numbers of atoms of all elements must be equal on both sides of the reaction arrow.

8. **List** three types of physical evidence that indicate a chemical reaction has occurred.

Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid.

9. **Compare and contrast** a skeleton equation and a chemical equation.

The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative amounts of reactants and products.

10. **Explain** why it is important to reduce coefficients in a balanced equation to the lowest possible whole-number ratio.

Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction.

11. **Analyze** When balancing a chemical equation, can you adjust the subscript in a formula? Explain.

No. Doing so changes the identity of the substance.

Chemistry Reactions Solutions Manual

**John Krenos, Joseph Potenza, Carl
Hoeger**



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Waldeck, Horst-Dieter Försterling, 2024-10-29 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text

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