



Singapore Junior Chemistry Olympiad

Mole Concept, Stoichiometry, and Redox

Full Notes

Learning Outcomes

You should be able to:

- (a) define the relative atomic, isotopic, molecular and formula masses, based on the ^{12}C scale,
- (b) define the mole in terms of the Avogadro constant,
- (c) calculate the relative atomic mass of an element given the relative abundances of its isotopes,
- (d) define the terms empirical and molecular formulae,
- (e) calculate empirical and molecular formulae, using combustion data or composition by mass,
- (f) write and/or construct balanced equations,
- (g) perform calculations, including use of the mole concept involving:
 - (i) reacting masses (from formulae and equations);
 - (ii) volumes of gases (e.g. in the burning of hydrocarbons);
 - (iii) volumes and concentrations of solutions,
- (h) deduce stoichiometric relationships from above calculations,
- (i) describe and explain redox processes in terms of electron transfer and/or of changes in oxidation number (oxidation state)
- (j) construct redox equations using the relevant half-equations.

Chemistry Olympiad Study Guide

RS Peters



Chemistry Olympiad Study Guide:

Decoding **Chemistry Olympiad Study Guide**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Chemistry Olympiad Study Guide**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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