



Chemistry

Paper Review and Study Guide

Name _____

Ionic & Metallic Compounds

1. Explain how the formation of ions contributes to the formation of ionic compounds.

2. Describe the characteristics of ionic compounds, including the strong electrostatic attraction between cations and anions.

Name _____

Ionic Compounds

21. How does the charge of ions affect the strength of ionic bonds?

22. Give an example of an ionic compound.

23. What are interstitial alloys?

24. Name a property unique to ionic compounds.

25. What type of bond is formed in metallic compounds?

26. What is the relationship between the strength of ionic bonds and the melting point of a compound?

28. How do you write the name for an ionic compound?

30. How do you write the formula for an ionic compound?

Name _____

Ionic Compounds & Metals

Study Guide:

1. Chemical Bonds

- Definition: The force that holds two atoms together.
- Types: ionic bonds, covalent bonds, and metallic bonds.

The formation of ions is a fundamental concept in chemistry, playing a crucial role in the formation of ionic compounds. Atoms achieve stability by attaining a full outer shell of electrons, resembling the electron configuration of noble gases. This process involves gaining or losing electrons, leading to the formation of ions. Metals, with their tendency to lose electrons, become positively charged ions called cations. Conversely, nonmetals gain electrons to become negatively charged ions known as anions.

2. Ion Formation

- Ions: An atom or molecule with a net electric charge due to the loss or gain of one or more electrons.
- Cations: A positively charged ion (loss of electrons).
- Anions: A negatively charged ion (gain of electrons).

Sodium (Na), a metal, readily loses its single valence electron to achieve the stable electron configuration of neon (Ne), resulting in a sodium cation (Na⁺). This cation now possesses 11 protons and 10 electrons, leading to a net positive charge. The formation of an ion is driven by the desire to achieve a stable electronic configuration, mimicking the noble gases, which are known for their chemical inertness.

3. Ionic Compounds

Ionic compounds, formed between oppositely charged ions, are the foundation of many materials. These bonds arise from the electrostatic attraction between cations and anions. The transfer of electrons from a metal atom to a non-metal atom results in the formation of these oppositely charged ions, which then combine to form a crystal lattice. The resulting ionic compounds exhibit characteristic properties, including high melting points and solubility in water.

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4. Alloy Formation

- Formation: Occurs between metals (which lose electrons) and non-metals (which gain electrons).
- Characteristics: Strong, directional bonds, leading to stable structures.



Ionic and Metallic Comps.

3-page Review +



39 Question Study Guide

Chemistry Metallic Bonds Study Guide Answers

**John Krenos, Joseph Potenza, Carl
Hoeger**



Chemistry Metallic Bonds Study Guide Answers:

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