

CHEMISTRY TEST

NAME: _____

INTRODUCTORY CHEMISTRY CONCEPTS

DATE: _____

Directions: For each of the following questions, choose the number that **best** answers the question and place it on your answer sheet.

Directions: For each of the following questions, all major steps must be shown to receive full credit.

Introductory Chemistry Concepts**Level 1****Multiple Choice**

1. A chemical change always results characteristically in a change of

- 1) state; 2) form; 3) identity; 4) size; 5) shape.

ANS: 3

2. Which of the following is a chemical change?

- 1) evaporation of water 2) distillation of water 3) freezing of water
4) electrolysis of water

ANS: 4

3. Chemical changes always involve 1) the absorption of energy; 2) the liberation of energy;
3) either the absorption or liberation of energy; 4) changes of state.

ANS: 3

4. One physical property of sugar is its 1) acidity; 2) reactivity;

- 3) flammability; 4) plasma qualities; 5) solubility.

ANS: 5

5. The majority of chemical changes are 1) autothermic; 2) endothermic;

- 3) isothermic; 4) nonthermic; 5) exothermic.

ANS: 5

6. The statement which describes a **physical** property of the element oxygen is

- 1) oxygen supports the burning of paper; 2) oxygen is needed for the rusting of iron;
3) oxygen must be present for food to decay; 4) oxygen has a density of 1.43 grams per liter;
5) oxygen is a component of water.

ANS: 4

7. In any chemical reaction energy is 1) always absorbed; 2) always released;

- 3) released as heat; 4) usually not involved; 5) either absorbed or released.

ANS: 5

8. Which of the following is **not** a chemical change?

- 1) the dissolving of zinc by HCl 2) the exploding of gasoline 3) the corroding of a copper roof
4) the boiling of water

ANS: 4

9. One chemical property of a substance is a description of its 1) odor; 2) density;

- 3) combustibility; 4) solubility; 5) none of these.

ANS: 3

10. A new substance is formed as a result of a 1) physical change; 2) chemical

Chemistry Matter And Change Studyguide Answer Key

H. Stephen Stoker, Danny V. White



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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