

Directions: The following is an End-Of-Course Review Guide designed to assist you as prepare for your EOC. It is imperative that you complete this guide to the best of your ability. This will help you to achieve a higher average on your second quarter grade.

Unit 1 & 2 Lab Safety, Introduction to Chemistry, & Measurement (Chapters 1 & 3):

1. What is chemistry?
2. Distinguish between hypothesis, theory, and law.
3. Classify the following as having good or poor accuracy and good or poor precision:
4. A scientist experimentally determines the speed of light to be 2.98×10^8 m/sec. In a second experiment, she determines the speed to be 2.99×10^8 m/sec.
5. The actual concentration of a solution is found to be 1.5 M. A scientist finds the concentrations to be 5.0 M and 5.2 M in two separate experiments.
6. Write the following in scientific notation:
a. 0.000 85 b. 1, 000, 000 c. 0.4565 d. 112
7. How many significant figures are in the following:
a. 0.98 b. 1, 000 c. 0.3450 d. 1.98×10^2
8. Report you answer to the correct number of significant figures:
a. $0.56 + 2.345$ b. $9.45 - 8.4$ c. 98.7×18.34 d. $1.23 / 3$
9. Explain how the measurement 45 cm is different from 45.0 cm.
10. List the SI units for the following quantities:
a. Mass b. Temperature c. Volume d. Length
11. How many meters are in the following:
a. 1 km b. 1 cm c. 1 mm
12. Calculate the density of a 5.0 g object that has a volume of 2.0 cm^3 .
13. Convert the following temperatures: a. 34°C to K b. 50 K to $^\circ\text{C}$
14. How will the graphs of a direct and indirect (inverse) proportion differ from each other?
15. Distinguish between periods and families.

Chemistry Eoc Study Guide

CO Houle



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