# CHAPTER 17 STUDY GUIDE

# **Chemical Equilibrium**

#### Section 17.1 Equilibrium: A State of Dynamic Balance

In your textbook, read about chemical equilibrium.

Complete each statement.

- When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to
- 2. A reaction that can occur in both the forward and the reverse directions is called a(n)

is a state in which the forward and reverse reactions balance each other because they take place at equal rates.

5. Equilibrium is a state of \_\_\_\_\_\_, not one of \_\_\_\_\_\_.

In your textbook, read about equilibrium expressions and constants.

For each statement below, write true or false.

- The law of chemical equilibrium states that at a given pressure, a chemical system
  may reach a state in which a particular ratio of reactant to product concentrations has
  a constant value.
- The equation H<sub>2</sub>(g) + I<sub>2</sub>(g) ←→ 2HI(g) is an example of a homogeneous equilibrium.
- If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.
- The value for K<sub>eq</sub> is constant only at a specific volume.
- If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.
- A heterogeneous equilibrium means that reactants and products are present in more than one state.
  - 12. The product of the forward chemical reaction is HI, for the equilibrium expression:

$$K_{eq} = \frac{[HI]^2}{[H_2][I_2]}$$

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**David R. Klein, Laurie S. Starkey** 

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