

General Chemistry II Chapter 14 Study Guide

Types of reaction rates:

- If the concentration of a reactant is plotted against time:
 - If the graph has a curved line = rate changes; will reach equilibrium
 - If the graph is linear = rate is constant; will run out of reactant
- Average rate of reaction: change in concentration of a reactant or product over time; reaction rate = slope
- instantaneous rate: slope of a line tangent to the curve; can pick any time
- Initial rate: rate within $t = 0$ to $t =$ given time; generally portion of the curve that is linear

Rate laws:

- Rate laws: only reactants appear in the rate law; the exponents tells the order of the rxn with respect to each reactant $\rightarrow$ sum of exponents is the overall order of the rxn; once we know the order of each reactant, we can plug in the concentrations and reaction rate to calculate the rate constant, k
- rate = $k[A]^m[B]^n$

Determining reaction orders:

- Suppose A and B are two reactants in a reaction:
 - If double $[A]$ will cause no change to rate, the order on A is 0 ($m=0$), e.g. $[2]0 = 1$, $[4]0 = 1$
 - If double $[A]$ will cause the rate to double, the order on A is 1 ($m=1$), e.g. $[2]1 = 2$, $[4]1 = 4$
 - If double $[A]$ will cause the rate to quadruple, the order on A is 2 ($m=2$), e.g. $[2]2 = 4$, $[4]2 = 16$
 - The order on reactant is not related to its coefficient in the chemical equation.
 - the exponents must be determined experimentally.

Determining rate laws: isolation method

- For rxns with more than one reactant, the rate law can be determined by the isolation method; approach:
 1. Fix the concentration of all reactants except for one $\rightarrow$ determine the rxn rates with varied concentrations of that reactant
 2. The change in rate is only attributed to that one reactant, enabling determination of the order of that reactant in the rate law
 3. Apply similar method to determine the order of other reactants

Zero-order process: overall order = 0:

Zero-Order Process: overall order = 0

For reaction $A \rightarrow \text{Product}$:

$$\text{Rate} = k[A]^0 = k$$

$$-\frac{d[A]}{dt} = k$$

$$d[A] = -k dt$$

Take the integral:

$$[A]_t = -kt + C$$

Plug in initial condition $[A] = [A]_0$ at $t=0$:

$$C = [A]_0$$

Units of k : $k = \text{rate} = \text{M/s}$



Zero-Order Process

Plot of $[A]_t$ vs. t



$$[A]_t = -kt + [A]_0$$

Equivalent to a straight-line function:
 $y = mx + b$

- Plot of $[A]_t$ vs. t is a straight line with $-k$ as slope.
- b is the y-intercept (i.e. the initial concentration $[A]_0$)
- Eventually $[A]$ will go to zero (all reactants will be used up)

Chemistry Chapter 14 Study Guide Answers

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Table of Contents Chemistry Chapter 14 Study Guide Answers

1. Understanding the eBook Chemistry Chapter 14 Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 14 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 14 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 14 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 14 Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 14 Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Chapter 14 Study Guide Answers Free and Paid eBooks
 - Chemistry Chapter 14 Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 14 Study Guide Answers eBook Subscription Services
 - Chemistry Chapter 14 Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 14 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 14 Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 14 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 14 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 14 Study Guide Answers
 - Interactive Elements Chemistry Chapter 14 Study Guide Answers
8. Staying Engaged with Chemistry Chapter 14 Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Chapter 14 Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 14 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 14 Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 14 Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 14 Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 14 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 14 Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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