

- A) They are restricted to metals.      C) Two reactants produce two products.  
 B) They involve a single product.      D) Any metal replaces any other metal.

16. State the type of reaction:  $2\text{H}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{O}_2 + \text{H}_2$

- A) synthesis      B) decomposition      C) single displacement  
 D) double displacement      E) combustion      F) no reaction

17. State the type of reaction:  $4\text{Cr} + 3\text{O}_2 \rightarrow 2\text{Cr}_2\text{O}_3$

- A) synthesis      B) decomposition      C) single displacement  
 D) double displacement      E) no reaction

18. Determine the number of moles present in 30.5-g aluminum chloride.

- A) 0.344 mol    B) 4.10 mol    C) 1.21 mol    D) 700 mol

19. What is the mass in grams of  $1.00 \times 10^{24}$  atoms of argon? (Ar)

- A)  $0.132 \times 10^3$  g    B)  $0.289 \times 10^3$  g    C)  $0.30 \times 10^3$  g    D)  $0.30 \times 10^3$  g

20. The number of molecules in 2.0 mol of carbon dioxide is:

- A)  $1.0 \times 10^{24}$     B)  $3.0 \times 10^{24}$     C)  $6.0 \times 10^{23}$     D)  $3.01 \times 10^{24}$     E)  $1.2 \times 10^{24}$

21. What is the percentage composition of COF?

- A) 50% C, 50% O    B) 12% C, 88% O    C) 25% C, 75% O    D) 40% C, 60% O

22. The mole ratio of carbon to hydrogen to oxygen in a compound is 1 mol C : 2 mol H : 1 mol O. What is the empirical formula of the compound?

- A) CHO    B) CH<sub>2</sub>O    C) C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>    D) C<sub>2</sub>H<sub>4</sub>O<sub>4</sub>

23. The molecular formula of a compound whose empirical formula is CH<sub>2</sub>O<sub>2</sub> and molar mass is 110 g/mol is:

- A) CH<sub>2</sub>O<sub>2</sub>    B) C<sub>2</sub>H<sub>4</sub>O<sub>4</sub>    C) C<sub>3</sub>H<sub>6</sub>O<sub>6</sub>    D) C<sub>4</sub>H<sub>8</sub>O<sub>8</sub>    E) C<sub>5</sub>H<sub>10</sub>O<sub>10</sub>

24. Lithium and nitrogen react in a combination reaction to produce lithium nitride:



In a particular experiment, 3.50-g samples of each reactant are reacted. The theoretical yield of lithium nitride is \_\_\_\_\_ g.

- A) 3.52    B) 2.89    C) 17.6    D) 5.85    E) 8.7

25. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:



The number of moles of KClO<sub>3</sub> required to produce 5.0 mol of oxygen is:

- A) 1.42    B) 0.422    C) 2.08    D) 4.38    E) 3.33

26. The theoretical yield of copper was calculated to be 288.0 g and the % yield was 80.90%. Determine the actual yield.

- A) 28.47 g    B) 264.3 g    C) 284.7 g    D) 205.1 g

# Chemistry 11 Final Exam Practice

**Minjie Lin**



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