



CONCENTRATION OF SOLUTIONS

The concentration of a solution is usually measured in moles per cubic decimetre (mol dm^{-3}). This is a measure of the number of moles in one cubic decimetre.

The volume must be in dm^3 (there are 1000 cm^3 in 1 dm^3). $\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$

$$\text{concentration (mol dm}^{-3}\text{)} = \frac{\text{moles}}{\text{volume (dm}^3\text{)}}$$



1) Calculate the concentration of the following solutions in mol dm^{-3} .

a) 0.10 moles of NaCl in 200 cm^3 _____

b) 0.20 moles of H_2SO_4 in 100 cm^3 _____

c) 0.020 moles of NaOH in 25 cm^3 _____

2) Calculate the number of moles in the following solutions.

a) 100 cm^3 of 0.20 mol dm^{-3} HNO_3 _____

b) 25 cm^3 of 1.50 mol dm^{-3} KOH _____

c) 50 cm^3 of 0.10 mol dm^{-3} H_2SO_4 _____

Concentration can also be measured in grams per cubic decimetre (g dm^{-3}). This is a measure of the number of grams in one cubic decimetre. (remember that mass = $A_r \times \text{moles}$)

1 dm^3

**2 moles of H_2SO_4 .
196 g of H_2SO_4 .**

Concentration = 2 mol dm^{-3}

MR of H_2SO_4 = 98

Concentration = 2 $\times$ 98 = 196 g dm^{-3}

A simple conversion is:

$$\text{conc (g dm}^{-3}\text{)} = \text{conc (mol dm}^{-3}\text{)} \times M_r$$

3) Calculate the concentration of the following solutions in g dm^{-3} .

a) 0.100 mol dm^{-3} NaOH _____

b) 0.250 mol dm^{-3} CH_3COOH _____

c) 1.50 mol dm^{-3} HNO_3 _____

4) 0.20 moles of NaOH is dissolved in 250 cm^3 of water.

a) Calculate the concentration in mol dm^{-3} _____

b) Calculate the concentration in g dm^{-3} _____

5) 5.0 g of KNO_3 is dissolved in 100 cm^3 of water.

a) Calculate the concentration in g dm^{-3} _____

b) Calculate the concentration in mol dm^{-3} _____

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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams/mole 21.3 grams / 0.66 moles = 32.4 g

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