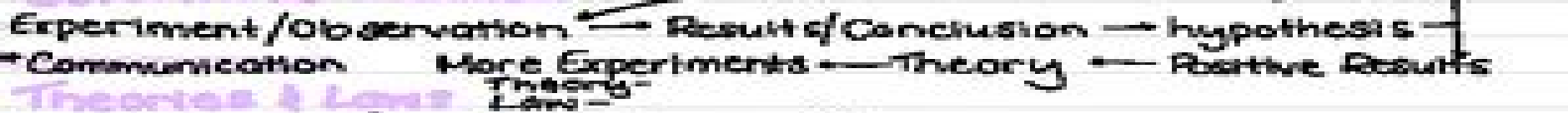


Finals Grinding

Chapter 1

1.1 Chemistry in Context

Scientific Method



1.2 Phases & Matter Classification

States of Matter

Solid - fixed shape and volume
 Liquid - shape of container, fixed volume
 gas - expands to fill container
 plasma - gaseous state w/ charged particles w/ unique properties

Mass & Weight

Mass - amt of matter
 Weight - force of gravity on object
 Conservation - Amt of matter doesn't change due to chemical or physical reactions

Defined Terms

Pure substance - constant composition
 Mixture - two or more substances
 Elements - same type of atom, smallest version
 Compounds - can be decomposed to other compounds/elements
 Homogeneous mixture - uniform composition
 Heterogeneous mixture - physically distinct parts

Miscible - forms a homogenous mixture when mixed

1.3 Physical & Chemical Properties

Ex/Intensive Properties

extensive depend on amt
 intensive independent of amt

Periodic Table

Atomic number = protons
 Periods - across
 Groups - down
 Chalcogens - group 16 (O, S, Se, Te, Po, Lv)
 alkali - alkali earth - transition ~ metals ~ metalloids ~ Non ~ Halo ~ Noble
 Chlorine - HCl
 Hydrobromic - HBr
 Nitric - HNO₃
 Hydrochloric - HCl
 Sulfuric - H₂SO₄
 Barium - Ba(OH)₂
 Calcium - Ca(OH)₂
 Strontium - Sr(OH)₂
 Lithium - LiOH
 Potassium - KOH
 Sodium - NaOH

1.4 Measurements

magnitude - size of number
 unit - standard of comparison
 indication of uncertainty - + or - uncertain digit
 Pico 10⁻¹² nano 10⁻⁹ micro 10⁻⁶ milli 10⁻³ centi 10⁻² deci 10⁻¹
 kilo 10³ mega 10⁶
 Density - mass/volume (invasive property)

1.5 Uncertainty, Accuracy, Precision

measurement device + user introduce uncertainty
 measure from bottom of meniscus

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