

Chemistry 33

The Ideal Gas Law

(Review for Chemistry 33)

Information:

- **Boyle's Law (1662)** – Volume or Molecular Occupancy $\propto 1/P$ (T is constant)
 - Boyle's Law says: $V \propto 1/P$ at constant T and n .
- **Charles's Law (1787)** – The volume of a sample of a gas varies directly with pressure, if the temperature is held constant.
 - Charles's Law says: $V \propto T$ at constant P and n .
- **Avogadro's Hypothesis (1811)** – Samples of different gases, which contain the same number of molecules (of any complexity, size, or shape) occupy the same volume at the same temperature and pressure.
 - Avogadro's Hypothesis says: $V \propto n$ at constant T and P .

Model 1: The Ideal Gas Law Equation.

$$PV = nRT$$

where P is pressure and V is the ideal gas constant.

The numerical value of the ideal gas constant is calculated from the fact that one mole of gas occupies 22.414 L at a pressure of one atmosphere and a temperature of 0°C (273.15 K):

$$R = \frac{PV}{nT} = \frac{(1 \text{ atm})(22.414 \text{ L})}{(1 \text{ mole})(273.15 \text{ K})} = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mole}}$$

Chemactivity 33 The Ideal Gas Law

Richard S. Moog, John J. Farrell



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Chemistry Richard S. Moog, John J. Farrell, 2014-01-13 Chemistry A Guided Approach 6th Edition follows the underlying principles developed by years of research on how readers learn and draws on testing by those using the POGIL methodology This text follows inquiry based learning and correspondingly emphasizes the underlying concepts and the reasoning behind the concepts This text offers an approach that follows modern cognitive learning principles by having readers learn how to create knowledge based on experimental data and how to test that knowledge **Ideal Gas Law, Enthalpy, Heat**

Capacity, Heats of Solution and Mixing Eric H Snider, 1984-01-01 Ideal Gas Law 51 Success Secrets - 51 Most Asked

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