

Stoichiometry Lab - **ANSWERS****PRE-LAB QUESTIONS:**

1. Calculate the molar mass of sodium bicarbonate (NaHCO_3). 84.003 g/mol
 2. Calculate the molar mass of acetic acid (CH_3COOH). 60.052 g/mol
 3. Calculate the molar mass of carbon dioxide (CO_2). 44.003 g/mol
 4. Vinegar is generally 4-5% acetic acid dissolved in distilled water. That means that a 5% vinegar solution contains 0.05 mol of acetic acid per liter of solution. How many moles of acetic acid are present in 30.00 mL?
- $$\frac{0.05 \text{ mol CH}_3\text{COOH}}{1 \text{ liter vinegar}} \times \frac{0.03 \text{ liter vinegar}}{1 \text{ liter vinegar}} = 0.0015 \text{ mol CH}_3\text{COOH in 30.0 mL of vinegar}$$
5. When baking soda (sodium bicarbonate) reacts with vinegar (acetic acid), they form sodium acetate, water, and carbon dioxide gas. Balance the chemical equation for this reaction.



6. A student measures out 5.00 g of baking soda to react with 50.00 mL of vinegar (which contains 0.042 moles of acetic acid).

a) Determine the limiting reagent, excess reagent, and theoretical yield (in grams) of carbon dioxide. Show your work.

$$5.00 \text{ g baking soda} \times \frac{1 \text{ mol baking soda}}{84.003 \text{ g baking soda}} = \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol baking soda}} \times \frac{5.00 \text{ g baking soda}}{1 \text{ mol baking soda}} = 0.06 \text{ g carbon dioxide}$$

$$0.042 \text{ moles acetic acid} \times \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol acetic acid}} \times \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 1.85 \text{ g carbon dioxide}$$

*** Since acetic acid produced the least amount of product, the acetic acid in vinegar is the **limiting reagent** and the baking soda is the **excess reagent**.

*** The theoretical yield of carbon dioxide is 1.85 g.

- b) This reaction will produce carbon dioxide bubbles that escape. After the reaction, the mass of the reaction decreased by 1.50 g. This is the actual mass of carbon dioxide that was produced from the reaction. Calculate the percentage yield of carbon dioxide. Show your work.

$$\text{Percent Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent Yield} = \frac{1.50 \text{ g}}{1.85 \text{ g}} \times 100 = 79.4\% \text{ yield}$$

Chemistry Stoichiometry Lab

Robert Bruce Thompson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2008-04-29 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier's Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry

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complement to a one semester preparatory chemistry laboratory course Tyner's manual emphasizes the application of chemistry and the principles of science to everyday life The labs are directly applicable to the real world and often contain supplemental assignments that illustrate an application

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The Journal of Industrial and Engineering Chemistry, 1910 **Electrochemical and Metallurgical Industry**, 1902 Sustainable Green Chemistry Mark Anthony Benvenuto, 2017-04-10 Sustainable Green Chemistry the 1st volume of Green Chemical Processing covers several key aspects of modern green processing The scope of this volume goes beyond bio and organic chemistry highlighting the ecological and economic benefits of enhanced sustainability in such diverse fields as petrochemistry metal production and wastewater treatment The authors discuss recent progresses and challenges in the implementation of green chemical processes as well as their transfer from academia to industry and teaching at all levels Selected successes in the greening of established processes and reactions are presented including the use of switchable polarity solvents actinide recovery using ionic liquids and the removal of the ubiquitous bisphenol A molecule from effluent streams by phytodegradation

Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

Research Based Undergraduate Science Teaching Dennis W. Sunal, Cynthia S Sunal, Emmett L. Wright, Cheryl L. Mason, Dean

Zollman,2014-07-01 Research in Science Education RISE Volume 6 Research Based Undergraduate Science Teaching examines research theory and practice concerning issues of teaching science with undergraduates This RISE volume addresses higher education faculty and all who teach entry level science The focus is on helping undergraduates develop a basic science literacy leading to scientific expertise RISE Volume 6 focuses on research based reforms leading to best practices in teaching undergraduates in science and engineering The goal of this volume is to provide a research foundation for the professional development of faculty teaching undergraduate science Such science instruction should have short and longterm impacts on student outcomes The goal was carried out through a series of events over several years The website at <http://nseus.org> documents materials from these events The international call for manuscripts for this volume requested the inclusion of major priorities and critical research areas methodological concerns and results of implementation of faculty professional development programs and reform in teaching in undergraduate science classrooms In developing research manuscripts to be reviewed for RISE Volume 6 researchers were asked to consider the status and effectiveness of current and experimental practices for reforming undergraduate science courses involving all undergraduates including groups of students who are not always well represented in STEM education To influence practice it is important to understand how researchbased practice is made and how it is implemented The volume should be considered as a first step in thinking through what reform in undergraduate science teaching might look like and how we help faculty to implement such reform

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 $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{H}_2 + \text{O}_2$
 automatically balanced by balancing nitrogen

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 b $\text{Mg}^{2+} + \text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{Mg}^{2+} + 2\text{NO}_3^- + \text{H}_2\text{O}$ write the balanced equation when nitric acid reacts with calcium hydroxide to produce water and calcium nitrate

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