Standardized Test prep Answers

Chapter 1 Page 25

1. C
2. C
3. D
4. B
5. C
6. C
7. C
8. A
9. B

10. Answers may vary.
11. Observations C and G are chemical properties; the others are physical properties.
12. In a chemical change, one or more substances are converted into different substances. A physical change does not involve a change in the identity of the substance or substances present.
13. metals: shiny; good conductors of heat; good conductors of electricity; malleable or ductile; most are solids at room temperature nonmetals: poor conductors of heat; poor conductors of electricity; many are gases at room temperature; those that are solids are brittle rather than malleable or ductile metalloids: properties intermediate between those of metals and nonmetals; less malleable than metals but not as brittle as solid nonmetals; most are semiconductors of electricity

Chapter 2 Page 63

1. B
2. C
3. C
4. B
5. A
6. B
7. A
8. B
9. B

10. 2 cups; direct proportion
11. Statements A and B contain exact numbers.
12. The type of fertilizer is the variable being tested. Control factors are the types of radishes, the amount of water and the amount of sunshine. One control row should be planted under the same control factors but with no fertilizer. There are at least four things that could be used to determine the results: size, quantity, appearance, and taste. Analysis might include bar graphs of each of these measurements for each of the five fertilizer types and the no-fertilizer control row.
13. A unit must be defined in a way that does not depend on the circumstances of the measurement. Not every thumbnail is the same size.

Chapter 3 Page 93

1. C
2. D
3. C
4. A
5. B
6. A
7. B
8. D
9. C

10. Argon-40 has 22 neutrons (40 − 18 = 22), and potassium-40 has 21 neutrons (40 − 19 = 21).
11. 6.17 g
12. All cathode rays are the same, regardless of their source. Therefore, the particles responsible for the cathode rays must be present in all atoms. The particles are electrons.
13. When the average atomic mass is calculated, it is 10.811. Because the atomic mass is the same as the atomic mass of boron, mythium was not a new element.

Chap 4 Page 129

1. A
2. D
3. A
4. D
5. B
6. B
7. C
8. B
9. C

10. The color will be orange. Converting energy into frequency gives 4.8×10^{14}, which corresponds to the frequency of orange light.
11. 

12. Electrons in atoms can occupy orbitals of only specific energies. When an atom is excited, the electron is no longer in the ground state. When the electron returns to a lower energy level, light is emitted. Because only specific energies are allowed, certain wavelengths of light are emitted, giving rise to the individual lines in the spectrum.
13. Photons of blue light are higher energy than photons of yellow light. Electrons can be emitted only when a photon of sufficient energy strikes the surface of the metal. Therefore, the energy of blue light is greater than the threshold energy, but the energy of yellow light is not.

Chapter 5 Page 1

1. C
2. A
3. A
4. A
5. C
6. A
7. D
8. B
9. C

10. S: 2251 kJ/mol; Cl: 2297 kJ/mol; Ar: 2666 kJ/mol; K: 3051 kJ/mol; Ca: 1145 kJ/mol; Sc: 1235 kJ/mol; Ti: 1310 kJ/mol. For the second ionization, the general trend is for increasing IE2 across the period in Groups 2–18 with Group 1 having the highest IE2. IE2 decreases going down a group.

11. Group 16 most commonly forms 2− ions, because these elements require only two more electrons to fill their shell (obtain a noble-gas configuration).

Chapter 6 Page 215

1. A
2. C
3. B
4. D
5. A
6. D
7. A
8. B
9. C

10. Hybridization explains how the orbitals of an atom become rearranged when the atom forms covalent bonds.
11. Ionic crystals are brittle because shifting of the layers of ions results in large repulsive forces that cause the layers to part completely.
12. Naphthalene has the stronger intermolecular forces even though it is non-polar, because its boiling point is higher than that of acetic acid. Boiling point is directly correlated to strength of intermolecular forces; the stronger the intermolecular forces, the more energy needed to break all the intermolecular forces, and therefore the higher the boiling point. Naphthalene is so large that its dispersion forces are greater than the sum of the dispersion forces and hydrogen bonding in acetic acid.
13. There is a repulsive region at small bond distances, a minimum well at an intermediate bond distance, and an asymptotic region at large bond distances. The repulsive region is due to the repulsion between the two positive nuclei at distances where the nuclei are close to each other. The asymptotic region is where the atoms are far apart and therefore do not interact with each other. The potential energy decreases as the atoms come closer together and are attracted to each other. The minimum well is the region where the attractive forces are ballanced by the repulsive forces. This distance is the equilibrium bond length.

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