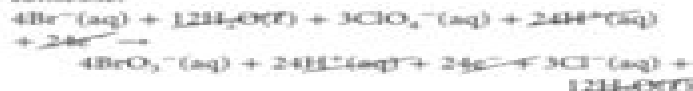


multiply reduction rx. by 3:



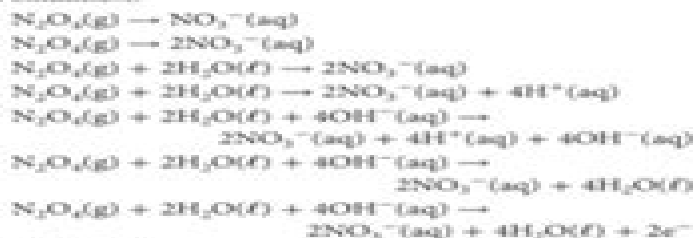
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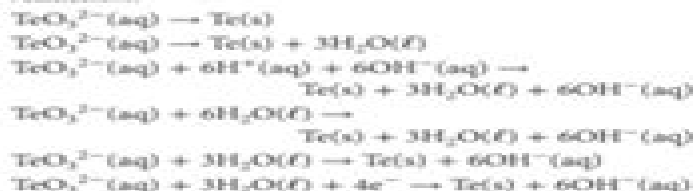
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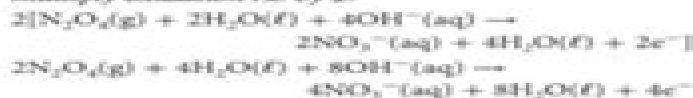
b. oxidation:



reduction:



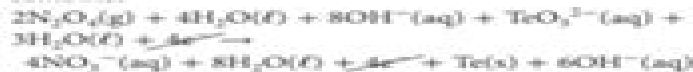
multiply oxidation rx. by 2:



multiply reduction rx. by 1:



combine:



balance:



c. oxidation: $\text{H}_2\text{O}_2(\text{aq}) \rightarrow \text{O}_2(\text{g})$



multiply oxidation rx. by 3:



multiply reduction rx. by 1:



combine:



14. It is a redox reaction because sulfur is both being oxidized from an oxidation number of -2 in $\text{H}_2\text{S}(\text{g})$ to an oxidation number of 0 in $\text{S}(\text{s})$ and reduced from an oxidation number of $+4$ in $\text{SO}_2(\text{g})$ to an oxidation number of 0 in $\text{S}(\text{s})$.

Answers to Section 9.3 Review Questions (Student textbook page 617)

- Fluorine is the most electronegative element in the periodic table. Therefore, it does not "lose" electrons in chemical bonds. It has seven valence electrons, and thus forms bonds in which it gains one valence electron, satisfying the octet rule. Because it is in the second row of the periodic table, its outer occupied energy level cannot accommodate more than eight electrons (as bromine, iodine, or chlorine can). Therefore, it does not form bonds in which it gains more than one electron, and its oxidation number is always -1 .
- $\text{MnO}_2 + 4$
 - $+7$
 - $+6$
 - $\text{KMnO}_4 + 7$
- Zero in the diatomic molecule, $\text{O}_2(\text{g})$; -1 in peroxides, $\text{H}_2\text{O}_2(\text{aq})$; -2 in most anions, $\text{CO}_3^{2-}(\text{aq})$; $+2$ in an oxygen difluoride, $\text{OF}_2(\text{g})$
- yes
 - no
 - no
 - yes

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A Loxley



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