

Name **KEY** Test Date Period **CP Chemistry – Final Exam Review***Show your work and round using significant figure rules for all the problems.*

1. What is chemistry?

- **Chemistry is the study of matter and the changes it undergoes.**

2. Write the following in scientific notation:

a. .005 **5×10^{-3}**

c. 53099000 **5.3099×10^7**

b. 5050 **5.05×10^3**

d. 0.00053030 **5.3030×10^{-4}**

3. Write the following in standard notation:

a. 1.5×10^3 **1500**

c. 4×10^{-5} **0.00004**

b. 3.75×10^{-2} **0.0375**

d. 9.12×10^7 **91200000**

4. How many significant figures are in the following:

a. 0.02 **1**

c. 501 **3**

b. 501.0 **4**

d. 0.020 **2**

5. Calculate and report in correct significant figures:

a. $52.568 \text{ cm} + 689.45 \text{ cm} + 55.5694 \text{ cm} =$ **797.59 cm**

b. $8426.200 \text{ g} \div 254 \text{ mol} =$ **33.2 g/mol**

6. You measure a Calcium cube to have 2.01 cm sides and a mass of 11.6 grams:

- a. What is the density as you calculate it?

$$V = s^3 = (2.01 \text{ cm})^3 = 8.12 \text{ cm}^3$$

$$d = \frac{m}{V} = \frac{11.6 \text{ g}}{8.12 \text{ cm}^3} = 1.43 \text{ g/cm}^3$$

- b. If calcium's accepted density is 1.54 g/cm
- ³
- , what is your percent error?

$$\% \text{error} = \frac{|1.54 \text{ g/cm}^3 - 1.43 \text{ g/cm}^3|}{1.54 \text{ g/cm}^3} \times 100 = 7.14 \%$$

7. Perform the following conversions:

- a. 53.1 mm to m

$$53.1 \text{ mm} \times \frac{1 \text{ m}}{10^3 \text{ mm}} = 0.0531 \text{ m}$$

- b. 18.4 kg to dag

$$18.4 \text{ kg} \times \frac{1 \times 10^{-1} \text{ dag}}{1 \times 10^{-3} \text{ kg}} = 1840 \text{ dag}$$

- c. 93.6 mi to km (1 mi = 1.6 km)

$$93.6 \text{ mi} \times \frac{1.6 \text{ km}}{1 \text{ mi}} = 150. \text{ km}$$

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David Floyd



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