

Vocab:

64. Unsaturated

65. Saturated
maximum amt. of solute dissolved at that temp.

66. Solubility
the amount (g) of solute that can dissolve at a certain temp.

67. Rate of dissolution
how fast a solute can dissolve

68. Molarity
concentration $M = \frac{\text{mol solute}}{1 \text{ Liter}}$

69. Supersaturated
more than max amt. of solute dissolved in solvent

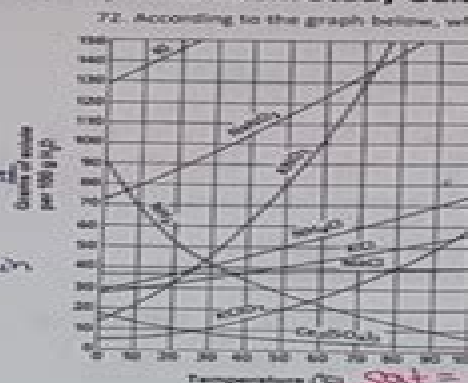
70. Electrolyte
- must be soluble
- must be ionic
can conduct electricity

71. List 3 ways to increase the rate of dissolution of a solid:

1. increase surface area (crush)
2. increase temperature
3. stir

SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER

Soluble compounds contain	Common exceptions
$\text{Cl}^-, \text{Br}^-, \text{I}^-, \text{NO}_3^-$	None
Ag^+	None
NO_3^-	None
Cl^-	None
CO_3^{2-}	None
OH^-	None
SO_4^{2-}	None
Br^-	None
I^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
NO_3^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
Cl^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
SO_4^{2-}	Compounds of Sr^{2+} , Ba^{2+} , Pb^{2+} , and Hg_2^{2+}
Insoluble compounds contain	Common exceptions
CO_3^{2-}	Compounds of Na^+ and the alkali metal cations
OH^-	Compounds of Na^+ and the alkali metal cations
ClO_4^-	Compounds of Na^+ and the alkali metal cations
Br^-	Compounds of Na^+ , the alkali metal cations, Ca^{2+} , Sr^{2+} , and Ba^{2+}
I^-	Compounds of Na^+ , the alkali metal cations, Ca^{2+} , Sr^{2+} , and Ba^{2+}



73. Which substance decreases in solubility with increasing temperature?

NH_3 , $\text{Ce}_2(\text{SO}_4)_3$

74. What is the solubility of KNO_3 at 60°C ?

$\sim 102\text{g}$

75. A solution containing 50 g CaCl_2 in 100 g H_2O is prepared at 10°C . Is the solution saturated, unsaturated or supersaturated?

(Not on graph :))

76. Use the solubility rules table on the left to help you with the following:

	Soluble	Insoluble	Dissociation reaction	Is it an electrolyte?	How many ions are produced?
NaNO_3	✓		$\text{NaNO}_3 \rightarrow \text{Na}^+ + \text{NO}_3^-$	✓	2
Na_2CO_3	✓		$\text{Na}_2\text{CO}_3 \rightarrow 2\text{Na}^+ + \text{CO}_3^{2-}$	✓	3
NaCl	✓		$\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$	✓	2
$\text{Ca}(\text{NO}_3)_2$	✓		$\text{Ca}(\text{NO}_3)_2 \rightarrow \text{Ca}^{2+} + 2\text{NO}_3^-$	✓	3
CaCO_3		✓	does not dissociate	NO	0

Molarity (M) is the number of moles of solute in one liter of solution

$$\text{molarity (M)} = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

Dilution: the action of making a stock solution less concentrated

$$M_1V_1 = M_2V_2$$

77. Calculate the Molarity of a solution that contains 100 grams of NaCl in 5.00L of solution [the molar mass of NaCl is 58.44 g/mol]

Convert g $\rightarrow$ mol

$$\frac{100 \text{ g NaCl}}{58.44 \text{ g/mol}} = 1.71 \text{ mol}$$

$$M = \frac{\text{mol}}{L}$$

$$M = \frac{1.71 \text{ mol}}{5.00 \text{ L}} =$$

0.34 molar solution

Page 5 of 7 **NaCl**

Chemistry Semester 2 Course Review Answers

ML Yell



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School and Society ,1923 *Directory of Distance Learning Opportunities* Modoc Press, Inc., 2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading

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