a matteres CH

Chemical Equation

Reactonts Left hand side.

-> Peroducts Right hand side

magnesium + Oxygen -> magnesium 0. - Ma0

BALANCING EQUATIONS

number of atoms of every element before and after meachion is same

$$H_2 + O_2 \longrightarrow H_2O$$
 $2H_2 + O_2 \longrightarrow 2H_2O$

Chemical equations are a way of showing chemical reactions.

LAW OF CONSErvation of matter a chemical equation must be balanced.

Chemical reactions

Chemical reactions can occurs only: wehen the chemical change. Two or more substance combine - Rusting in Iron . Burning of coal . to form a single product.



DECOMPOSITION -

Single meachint breaks up to form tow or more products

→ 2H.O --> 2H. + O.

DISPLACEMENT < more reactive element removes the other element form the compound and itself forms the compound

>> Pb+ Cacl2 --> PbCl2 + Ca

DOUBLE DISPLACEMENT

Exchange of ions between the reactants take place.

W NaOH + HCI -- NaCI+HO

REDUCTION -

gain of hydrogen or loss of oxygen.

gain of oxygen

or loss of hydrogen. oxidation

- Cu0 + H2 - Cu + H-0 reduction destudy 102

sedox «

2Pb(NO₃),(s) heat > 2PbO(s) + 4NO,(s) + 0,(s) PHYSICAL STATE -

(cs) + 02 cg) -- (0, cg) + Q Exothermic CHANGE IN HEAT -

N2 (g) + O2 (g) - 2NO (g) - Q Endothermic

Zncs) + H,504(aq) --- > ZnS04(aq) + H,1 EVOLUTION OF GOS -

AgNO, cag) + Na(1 cag) --- Ag(1 + NaNo, (ag) PRECIPITATION -

DIFF. REACTIONS CONDITION - GCOZ (9) + GHZO(1) SUNTIGHT (6 HIZ Oc caq) + 60269)

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Matt Kingsley

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