

Chemistry Study Guide: Unit 6: Gases

Kinetic Molecular Theory (KMT): a model or theory that is used to explain the behavior of gases.

- A. Gases contain particles (molecules or atoms) that are in constant, random, straight-line motion.
- B. Gas particles collide with each other and with the walls of the container. These collisions are perfectly elastic, meaning kinetic energy (energy an object has because of its motion) is being transferred among the particles but no kinetic energy is lost.
- C. Gas particles are separated by relatively great distances. Because of this, the volume of the particles themselves is negligible (insignificant).
- D. Gas particles do not attract or repel each other. This is because they are far away from each other.

Gas Pressure: Force exerted by a gas on the walls of its container.

- Caused by collisions between gas particles and walls of the container.
- Units: kilopascal (kPa), Atmosphere (atm), millimeters mercury (mm Hg)
- $1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$
- Found on table A with temperature on periodic table

Temperature: a measure of the average kinetic energy of the particles in a substance.

- As average kinetic energy of particles in a substance increases, temperature of a substance increases.
- In science, Celsius and kelvin are common temperature scales.
- $K = C^{\circ} + 273$, $C^{\circ} = K - 273$
- Temperature is not a form of energy.

Relationship between pressure and volume of a gas:

Boyle's Law: Pressure and Volume (Temperature Constant)

If the temperature is constant, then the pressure of a gas increases, as the volume decreases. As the pressure decreases the volume increases.

- Inverse relationship between pressure and volume.

Equation: $P_1 \bullet V_1 = P_2 \bullet V_2$

Charles Law: Temperature and Volume (Pressure Constant)

If the pressure is constant, then as the temperature of an enclosed gas increases, the volume increases. As the temperature decreases, the volume decreases.

- Direct relationship between temperature and volume
- Temperature must be converted to Kelvin.

Equation: $V_1 / T_1 = V_2 / T_2$

Gay-Lussac's Law: Temperature and Pressure (Volume is constant)

If the volume is constant, then as the temperature of an enclosed gas increases, the pressure increases. As the temperature decreases, the pressure decreases.

- Direct relationship between temperature and pressure.
- Temperature must be converted to Kelvin.

Equation: $P_1 / T_1 = P_2 / T_2$

Chemistry Gases Unit Study Guide

Jill Meryl Levy



Chemistry Gases Unit Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004. The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even-numbered end-of-chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application.

Organic Chemistry, Student Study Guide and Solutions Manual David R. Klein, 2017-01-04. This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e. Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis. Success in organic chemistry requires mastery in two core aspects: fundamental concepts and the skills needed to apply those concepts and solve problems. Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills. These skills are vital for successful problem solving in organic chemistry. Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems.

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Table of Contents Chemistry Gases Unit Study Guide

1. Understanding the eBook Chemistry Gases Unit Study Guide
 - The Rise of Digital Reading Chemistry Gases Unit Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Gases Unit Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Gases Unit Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Gases Unit Study Guide
 - Personalized Recommendations
 - Chemistry Gases Unit Study Guide User Reviews and Ratings
 - Chemistry Gases Unit Study Guide and Bestseller Lists
5. Accessing Chemistry Gases Unit Study Guide Free and Paid eBooks
 - Chemistry Gases Unit Study Guide Public Domain eBooks
 - Chemistry Gases Unit Study Guide eBook Subscription Services
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6. Navigating Chemistry Gases Unit Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Gases Unit Study Guide Compatibility with Devices
 - Chemistry Gases Unit Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Gases Unit Study Guide
 - Highlighting and Note-Taking Chemistry Gases Unit Study Guide
 - Interactive Elements Chemistry Gases Unit Study Guide
8. Staying Engaged with Chemistry Gases Unit Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Gases Unit Study Guide
- 9. Balancing eBooks and Physical Books Chemistry Gases Unit Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Gases Unit Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Gases Unit Study Guide
 - Setting Reading Goals Chemistry Gases Unit Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Gases Unit Study Guide
 - Fact-Checking eBook Content of Chemistry Gases Unit Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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