



# CHEMISTRY

## CHEAT SHEET



# NAMING SIMPLE COMPOUNDS

## NAMING SIMPLE COMPOUNDS CHEAT SHEET

Naming Compounds with Their Formulas  
(Show the name corresponding to the formula)



**Directions for naming:**

1. Name the cation first, then the anion.
2. If the cation is a transition metal, include its oxidation state in parentheses.
3. If the anion is a polyatomic ion, use its name.
4. If the cation is a transition metal, use the suffix -ous or -ic.
5. If the anion is a polyatomic ion, use the suffix -ate or -ite.
6. If the compound is a binary compound, use the suffix -ide.

| Binary Compounds                |                    | Common Names (Binary Compounds) |                    |
|---------------------------------|--------------------|---------------------------------|--------------------|
| Formula                         | Name               | Formula                         | Name               |
| NaCl                            | sodium chloride    | NaCl                            | sodium chloride    |
| NaBr                            | sodium bromide     | NaBr                            | sodium bromide     |
| NaI                             | sodium iodide      | NaI                             | sodium iodide      |
| NaF                             | sodium fluoride    | NaF                             | sodium fluoride    |
| Na <sub>2</sub> O               | sodium oxide       | Na <sub>2</sub> O               | sodium oxide       |
| Na <sub>2</sub> S               | sodium sulfide     | Na <sub>2</sub> S               | sodium sulfide     |
| Na <sub>2</sub> SO <sub>4</sub> | sodium sulfate     | Na <sub>2</sub> SO <sub>4</sub> | sodium sulfate     |
| Na <sub>2</sub> CO <sub>3</sub> | sodium carbonate   | Na <sub>2</sub> CO <sub>3</sub> | sodium carbonate   |
| NaHCO <sub>3</sub>              | sodium bicarbonate | NaHCO <sub>3</sub>              | sodium bicarbonate |
| NaOH                            | sodium hydroxide   | NaOH                            | sodium hydroxide   |

Naming Compounds with a Central and Terminal



**Carbon Dioxide**

**Directions for naming:**

1. Name the cation first, then the anion.
2. If the cation is a transition metal, include its oxidation state in parentheses.
3. If the anion is a polyatomic ion, use its name.
4. If the cation is a transition metal, use the suffix -ous or -ic.
5. If the anion is a polyatomic ion, use the suffix -ate or -ite.
6. If the compound is a binary compound, use the suffix -ide.

**Common Names**

| Formula                         | Common Name        | Formula                         | Common Name        |
|---------------------------------|--------------------|---------------------------------|--------------------|
| NaCl                            | sodium chloride    | NaCl                            | sodium chloride    |
| NaBr                            | sodium bromide     | NaBr                            | sodium bromide     |
| NaI                             | sodium iodide      | NaI                             | sodium iodide      |
| NaF                             | sodium fluoride    | NaF                             | sodium fluoride    |
| Na <sub>2</sub> O               | sodium oxide       | Na <sub>2</sub> O               | sodium oxide       |
| Na <sub>2</sub> S               | sodium sulfide     | Na <sub>2</sub> S               | sodium sulfide     |
| Na <sub>2</sub> SO <sub>4</sub> | sodium sulfate     | Na <sub>2</sub> SO <sub>4</sub> | sodium sulfate     |
| Na <sub>2</sub> CO <sub>3</sub> | sodium carbonate   | Na <sub>2</sub> CO <sub>3</sub> | sodium carbonate   |
| NaHCO <sub>3</sub>              | sodium bicarbonate | NaHCO <sub>3</sub>              | sodium bicarbonate |
| NaOH                            | sodium hydroxide   | NaOH                            | sodium hydroxide   |

**Common Names for Specific Compounds**

| Formula                         | Common Name        | Formula                         | Common Name        |
|---------------------------------|--------------------|---------------------------------|--------------------|
| NaCl                            | sodium chloride    | NaCl                            | sodium chloride    |
| NaBr                            | sodium bromide     | NaBr                            | sodium bromide     |
| NaI                             | sodium iodide      | NaI                             | sodium iodide      |
| NaF                             | sodium fluoride    | NaF                             | sodium fluoride    |
| Na <sub>2</sub> O               | sodium oxide       | Na <sub>2</sub> O               | sodium oxide       |
| Na <sub>2</sub> S               | sodium sulfide     | Na <sub>2</sub> S               | sodium sulfide     |
| Na <sub>2</sub> SO <sub>4</sub> | sodium sulfate     | Na <sub>2</sub> SO <sub>4</sub> | sodium sulfate     |
| Na <sub>2</sub> CO <sub>3</sub> | sodium carbonate   | Na <sub>2</sub> CO <sub>3</sub> | sodium carbonate   |
| NaHCO <sub>3</sub>              | sodium bicarbonate | NaHCO <sub>3</sub>              | sodium bicarbonate |
| NaOH                            | sodium hydroxide   | NaOH                            | sodium hydroxide   |

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**SA Dillow**



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