

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + 2 \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + 2 \text{NaBr} \rightarrow \text{MgBr}_2 + 2 \text{Na}$
- $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- $4 \text{P} + 5 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + 6 \text{KOH} \rightarrow 3 \text{K}_2\text{SO}_4 + 2 \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow \text{Fe} + 3 \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on right, 2 on left

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on right, 2 on left

12. Using the balanced equation shown in problem #10, list the products of the reaction.

NH_3

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N_2 , H_2

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



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Robert R. Redfield



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