

EQUILIBRIUM CHEAT SHEET

Chemical equilibrium, dynamic equilibrium, when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium reactions are represented by this symbol: $\rightleftharpoons$

What is Equilibrium?

Equilibrium is a constant concentration of reactants and products.

When a reversible reaction takes place in a closed system, the forward and reverse reactions continue to occur. When a system is at equilibrium, the concentrations of reactants and products do not change.

It is interesting to think of equilibrium as a constant state, but the forward and reverse reactions are still occurring. This is called a dynamic equilibrium, as the system is still changing, but the concentrations of reactants and products are constant.

Equilibrium Constant Formed

$$aA + bB \rightleftharpoons cC + dD$$

Equilibrium Constant, K_c

When a reversible reaction is at equilibrium, the ratio of the concentrations of products to reactants is constant. This is called the equilibrium constant, K_c . It is a measure of the extent to which a reaction proceeds.

$$aA + bB \rightleftharpoons cC + dD$$

Equilibrium Constant Equation

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

The equilibrium constant is a constant for a given reaction at a given temperature. It is a measure of the extent to which a reaction proceeds.

Reaction Quotient, Q

The reaction quotient, Q , is a measure of the extent to which a reaction has proceeded. It is calculated using the same formula as the equilibrium constant, K_c , but with the concentrations of reactants and products at any point in time, rather than at equilibrium.

$$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constant

$$aA + bB \rightleftharpoons cC + dD$$

Example



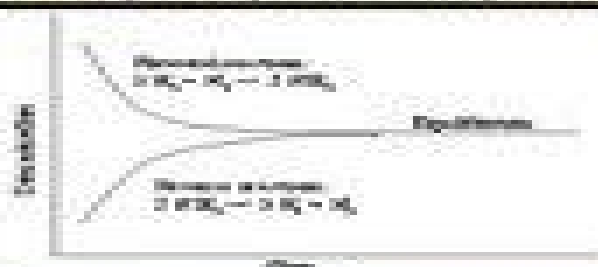
Forward Reaction



Reverse Reaction



Graphical Representation of Equilibrium



Equilibrium Constant Relationship

Equilibrium Constant Value	Meaning
$K_c > 1$	Concentration of products is greater than concentration of reactants
$K_c < 1$	Concentration of reactants is greater than concentration of products
$K_c = 1$	Concentration of reactants is equal to concentration of products

Reaction Quotient vs. Equilibrium Constant

Equilibrium Constant vs. Reaction Quotient	Meaning
$Q < K_c$	Reaction is not at equilibrium. Forward reaction is favored. $2H_2 + O_2 \rightarrow 2H_2O$
$Q > K_c$	Reaction is not at equilibrium. Reverse reaction is favored. $2H_2O \rightarrow 2H_2 + O_2$
$Q = K_c$	Reaction is at equilibrium. $2H_2 + O_2 \rightleftharpoons 2H_2O$

Le Chatelier's Principle: The effect of changing conditions on equilibrium

Condition	Effect of Change	Example
Concentration	- Increasing concentration of reactants shifts equilibrium to the right. - Increasing concentration of products shifts equilibrium to the left.	$2H_2 + O_2 \rightleftharpoons 2H_2O$ Increasing concentration of H_2 or O_2 shifts equilibrium to the right, increasing concentration of H_2O .
Pressure	When pressure changes, the system shifts to the side with fewer moles of gas. - Increasing pressure shifts equilibrium to the side with fewer moles of gas. - Decreasing pressure shifts equilibrium to the side with more moles of gas.	$2H_2 + O_2 \rightleftharpoons 2H_2O$ Increasing pressure shifts equilibrium to the right, as there are fewer moles of gas on the right side.
Temperature	When temperature changes, the system shifts to the side that absorbs heat. - Increasing temperature shifts equilibrium to the endothermic side. - Decreasing temperature shifts equilibrium to the exothermic side.	$2H_2 + O_2 \rightleftharpoons 2H_2O$ Increasing temperature shifts equilibrium to the left, as the forward reaction is exothermic.

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