

pH AND pOH

Chapter 14 Section 14.2

The pH of a solution indicates how acidic or basic that solution is.

pH ranges of 0-7 = acidic

7 = neutral

7-14 = basic

Since $pH + pOH = 14$ for all aqueous solutions, then pOH can be calculated using either value.

$pH = 4$ implies

So if $pH = 4$ then $pOH = 10$

$pOH = 10$ implies

So if $pOH = 10$ then $pH = 4$

Together $pH + pOH = 14$

Complete the following chart.

	pH	pH	pOH	pOH	Acids or Bases
1.	normal	7	normal	7	Neutral
2.	10^{-12}	12	10^{-2}	2	acidic
3.	10^{-10}	10	normal	4	acidic
4.	normal	7	10^{-7}	7	neutral
5.	10^{-7}	7	10^{-7}	7	neutral
6.	10^{-11}	11	10^{-3}	3	acidic
7.	10^{-9}	9	normal	5	acidic
8.	10^{-6}	6	normal	8	basic
9.	10^{-4}	4	10^{-10}	10	basic
10.	10^{-2}	2	10^{-12}	12	basic

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