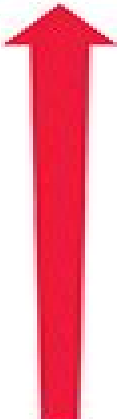
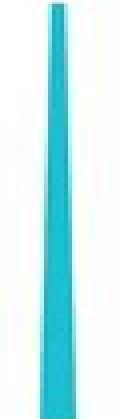
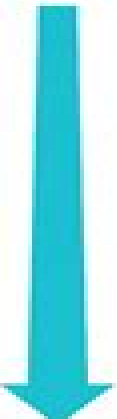


Acids and Bases

MERCK

Acid				Base		
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base
Strongest 	Large	Perchloric acid	HClO ₄	ClO ₄ ⁻	Perchlorate ion	Weakest 
	3.2×10 ⁸	Hydroiodic acid		I ⁻	Iodide	
	1.0×10 ⁹	Hydrobromic acid	HBr	Br ⁻	Bromide	
	1.3×10 ⁹	Hydrochloric acid	HCl	Cl ⁻	Chloride	
	1.0×10 ³	Sulfuric acid	H ₂ SO ₄	HSO ₄ ⁻	Hydrogen sulfate ion	
	2.4×10 ¹	Nitric acid	HNO ₃	NO ₃ ⁻	Nitrate ion	
	-----	Hydronium ion	H ₃ O ⁺	H ₂ O	Water	
	5.4×10 ⁻²	Oxalic acid	HO ₂ C ₂ O ₂ H	HO ₂ C ₂ O ₂ ⁻	Hydrogen oxalate ion	
	1.3×10 ⁻²	Sulfurous acid	H ₂ SO ₃	HSO ₃ ⁻	Hydrogen sulfite ion	
	1.0×10 ⁻²	Hydrogen sulfate ion	HSO ₄ ⁻	SO ₄ ²⁻	Sulfate ion	
Medium 	7.1×10 ⁻³	Phosphoric acid	H ₃ PO ₄	H ₂ PO ₄ ⁻	Dihydrogen phosphate ion	Medium 
	7.2×10 ⁻⁴	Nitrous acid	HNO ₂	NO ₂ ⁻	Nitrite ion	
	6.6×10 ⁻⁴	Hydrofluoric acid	HF	F ⁻	Fluoride ion	
	1.8×10 ⁻⁴	Methanoic acid	HCO ₂ H	HCO ₂ ⁻	Methanoate ion	
	6.3×10 ⁻⁵	Benzoic acid	C ₆ H ₅ COOH	C ₆ H ₅ COO ⁻	Benzoate ion	
	5.4×10 ⁻⁵	Hydrogen oxalate ion	HO ₂ C ₂ O ₂ ⁻	O ₂ C ₂ O ₂ ²⁻	Oxalate ion	
	1.8×10 ⁻⁵	Ethanoic acid	CH ₃ COOH	CH ₃ COO ⁻	Ethanoate (acetate) ion	
	4.4×10 ⁻⁷	Carbonic acid	H ₂ CO ₃	HCO ₃ ⁻	Hydrogen carbonate ion	
	1.1×10 ⁻⁷	Hydrosulfuric acid	H ₂ S	HS ⁻	Hydrogen sulfide ion	
	6.3×10 ⁻⁸	Dihydrogen phosphate ion	H ₂ PO ₄ ⁻	HPO ₄ ²⁻	Hydrogen phosphate ion	
Weakest	6.2×10 ⁻⁸	Hydrogen sulfite ion	HS ⁻	SO ₃ ²⁻	Sulfite ion	Strongest
	2.9×10 ⁻⁸	Hypochlorous acid	HClO	ClO ⁻	Hypochlorite ion	
	6.2×10 ⁻¹⁰	Hydrocyanic acid	HCN	CN ⁻	Cyanide ion	
	5.8×10 ⁻¹⁰	Ammonium ion	NH ₄ ⁺	NH ₃	Ammonia	
	5.8×10 ⁻¹⁰	Boric acid	H ₃ BO ₃	H ₂ BO ₃ ⁻	Dihydrogen borate	
	4.7×10 ⁻¹¹	Hydrogen carbonate ion	HCO ₃ ⁻	CO ₃ ²⁻	Carbonate ion	
	4.2×10 ⁻¹³	Hydrogen phosphate ion	HPO ₄ ²⁻	PO ₄ ³⁻	Phosphate ion	
	1.8×10 ⁻¹³	Dihydrogen borate ion	H ₂ BO ₃ ⁻	HBO ₃ ²⁻	Hydrogen borate ion	
	1.3×10 ⁻¹³	Hydrogen sulfide ion	HS ⁻	S ²⁻	Sulfide ion	
	1.6×10 ⁻¹⁴	Hydrogen borate ion	HBO ₃ ²⁻	BO ₃ ³⁻	Borate ion	
	-----	water	H ₂ O	OH ⁻	Hydroxide	



Chemistry Acids And Bases Guide

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



Chemistry Acids And Bases Guide:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual* T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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Survival Guide to Organic Chemistry Bridging the Gap from General Chemistry enables organic chemistry students to bridge the gap between general chemistry and organic chemistry It makes sense of the myriad of in depth concepts of organic chemistry without overwhelming them in the necessary detail often given in a complete organic chemistry text Here the topics covered span the entire standard organic chemistry curriculum The authors describe subjects which require further explanation offer alternate viewpoints for understanding and provide hands on practical problems and solutions to help master the material This text ultimately allows students to apply key ideas from their general chemistry curriculum to key concepts in organic chemistry Key Features Reviews key general chemistry concepts and techniques adapted for application to important organic principles Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry Explains organic concepts and reaction mechanisms generally expanding the focus on how to understand each step from a more intuitive viewpoint Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms Implements a user friendly process to achieve the end result of problem solving Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum features practice problems and solutions to help master the material including an extensive and comprehensive bank of practice exams with solutions

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