HSC Chemistry Notes:

- Module 5: Equilibrium and Acid Reactions
- Module 6: Acid and Base Reactions
- Module 7: Organic Chemistry
- Module 8: Applying Chemical Ideas

Module 5: Equilibrium and Acid Reactions

Static and Dynamic Equilibrium

Reversibility of chemical reactions:

Formation of hydrated and dehydrated cobalt (II) chloride:

Dehydrated cobalt (II) chloride + water ⇌ hydrated cobalt (II) chloride

$$CoCl_2 + 6H_2O \rightleftharpoons CoCl_2 \cdot 6H_2O$$

Blue pink

- When the hydrated cobalt (II) chloride is heated, it becomes blue again as the water evaporated.

Reaction of Iron (III) nitrate and potassium thiocyanate:

Iron (III) nitrate + potassium thiocyanate ⇌ iron (iii) thiocyanate

$$Fe^{3+}_{(aq)} + SCN^-_{(aq)} \rightleftharpoons FeSCN^{2+}_{(aq)}$$

Pale yellow colourless red

- When ($Fe(NO_3)_3$) and (KSCN) are mixed, they form a red-coloured ion, iron (III) thiocyanate
- Is a reversible reaction and forms an equilibrium

Burning Magnesium:

$$2Mg_{(s)} + O_2 \rightarrow 2MgO_{(s)}$$

- An irreversible combustion reaction, releasing energy into the system, producing white magnesium oxide.

Burning Steel Wool:

$$2Fe_{(s)} + \frac{3}{2}O_{2(g)} \rightarrow Fe_2O_{3(s)}$$

- An irreversible reaction and is a combustion reaction.
- As the steel wool caught on fire and went red with sparks flying out.

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CH Cherryholmes

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