

CHE 101L Lab 7 Chemicals in Everyday Life Post Lab Questions

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Dept. of Natural and Physical Sciences

1. How could you distinguish sodium chloride (table salt) from sodium iodide? Show reactions equations.

Sodium chloride can be distinguished from sodium iodide by looking at the precipitate formed after adding the salts to separate silver nitrate solutions. Sodium chloride reacts with silver nitrate to form a white silver chloride precipitate, while sodium iodide reacts with silver nitrate to form a yellow silver iodide precipitate. Another way to distinguish sodium chloride from sodium iodide is by adding the salts to concentrated sulfuric acid. Sodium chloride forms a colorless hydrogen chloride gas, while sodium iodide forms hydrogen sulfide gas and a brown tri iodide solid.

Add copper sulfate to both:



This would appear brown due to the formation of I_2 .

CuSO_4 plus NaCl would perhaps turn slightly green owing to formation of some CuCl_2 .

Add AgNO_3 to both.

The resulting pptes have different colors



2. How could you distinguish solid barium chloride from barium sulfate?

They can be distinguished by using water. Barium chloride is soluble in water while Barium sulfate is insoluble in water but soluble in methanol. Also, BaCl_2 forms crystals; while BaSO_4 is a morph and does not form crystals.

3. Do you think that washing soda Na_2CO_3 could be used for the same purpose as baking soda NaHCO_3 ? Would Na_2CO_3 react with HCl ? Write the chemical equations for the reaction of HCl with both,

No, it cannot be used because washing soda will consume two equivalents of acid, while baking soda will only consume one equivalent. Basically, Washing Soda is a stronger base than baking soda, and is in fact, CAUSTIC. This is one reason why it isn't used for baking. NaHCO_3 will decompose to Na_2CO_3 , CO_2 and H_2O at high temperatures. That is why it is used as a leavening

Chemicals In Everyday Life Lab Answers

Vânia Zuin, Lílana Mammino



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