Chemical Reactions and Equations

A. Magnesium and Oxygen

1. Initial appearance of Mg Shiny light gray ribbon of metal

2. Evidence of chemical reaction Heat and light energy produced

White ash produced

3. Balance: $2 \text{ Mg}_{00} + O_{Ng} \stackrel{1}{\rightarrow} 2 \text{ Mg}O_{00}$

4. Type of chemical reaction: Combination reaction

B. Zinc and Copper(II) Sulfate

| Time | CuSO ₄ (aq) Appearance | | Evidence of a Chemical Reaction |
|--------------------|--------------------------------------|-----------|------------------------------------|
| 1. initial | Blue/green | Dull gray | |
| 2. after 30 min | Lt blue | Black | Change in color |

3. Zn₆₀ + CuSO₄₀₄₀

¹ Cu₆₀ + ZnSO₄₀₄₀ (the equation is balanced)

4. Type of chemical reaction: Single replacement reaction

C. Reactions of Metals and HC1

| Metal | 1. Appearance of Metals | 2. Evidence of a Chemical Reaction |
|-------|--------------------------|---|
| Zn(s) | Gray, shiny pebbles | Formation of bubbling |
| Cu(s) | Orange, shiny strips | No change |
| Mg(s) | Light gray, shiny strips | Formation of many bubble Heat produced |

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