

NAME: _____ Period: _____

Ch 10 Chemical Quantities Study Guide

/75

Matching

Match each item with the correct statement below.

- a. molar volume
- b. molar mass
- c. atomic mass

- _____ 1. the number of grams of an element that is numerically equal to the atomic mass of the element in amu
- _____ 2. the mass of a mole of any element or compound
- _____ 3. the volume occupied by a mole of any gas at STP

Match each item with the correct statement below.

- a. representative particle
- b. mole
- c. Avogadro's number
- d. percent composition
- e. standard temperature and pressure
- f. empirical formula

- _____ 4. the number of representative particles of a substance present in 1 mole of that substance
- _____ 5. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- _____ 6. the SI unit used to measure amount of substance
- _____ 7. 0°C and 1 atm
- _____ 8. the percent by mass of each element in a compound
- _____ 9. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. What SI unit is used to measure the number of representative particles in a substance?
 - a. kilogram
 - b. ampere
 - c. kelvin
 - d. mole
- _____ 2. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C_3H_7O ?
 - a. $5 \times (6.02 \times 10^{23})$
 - b. 5
 - c. 35
 - d. $35 \times (6.02 \times 10^{23})$
- _____ 3. If a dozen Clementine oranges have a mass of 744 g, what will be the mass of 15 Clementines?
 - a. 11.16 kg
 - b. 595 g
 - c. 896 g
 - d. 930 g
- _____ 4. Which of the following elements exists as a diatomic molecule?
 - a. neon
 - b. lithium
 - c. nitrogen
 - d. sulfur
- _____ 5. Avogadro's number of representative particles is equal to one _____.
 - a. kilogram
 - b. gram
 - c. kelvin
 - d. mole
- _____ 6. Which of the following is equal to Avogadro's number?
 - a. the number of atoms of bromine in 1 mol Br_2
 - b. the number of molecules of chlorine in 0.50 mol Cl_2
 - c. the number of molecules of nitrogen in 1 mol N_2
 - d. the number of atoms in 1 mol CO

Chemical Quantities Study Guide Chemistry Answers

Iris Stovall



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement **Study Guide for Introductory Chemistry** Iris Stovall, 1990

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study guide will be of great value to chemistry teachers and students **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry Iris Stovall, Steven S. Zumdahl, 1996 **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall, Donald J. DeCoste, 2000

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