



$$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b} \quad (2a)$$

$$K_{eq} = \frac{[C]_{eq}^c [D]_{eq}^d}{[A]_{eq}^a [B]_{eq}^b} \quad (2b)$$

Reaction Quotient

Equilibrium constant

$$\Delta G = \Delta G^\circ + RT \ln Q = \Delta G^\circ + RT \ln \left(\frac{[C]^c [D]^d}{[A]^a [B]^b} \right) \quad (3)$$

$$\Delta G = 0 \quad \text{and} \quad Q = K_{eq} \quad (4)$$

Equilibrium conditions

$$\Delta G^\circ = -RT \ln K_{eq} = -\ln 10 \cdot RT \log_{10} K_{eq} \quad (5)$$

$$K_{eq} = \exp(-\Delta G^\circ / RT) = 10^{-\Delta G^\circ / (2.3026 RT)} \quad (6)$$

Chemical Engineering Thermodynamics The Study Of Energy Entropy Equilibrium

M Carnoy



Chemical Engineering Thermodynamics The Study Of Energy Entropy Equilibrium:

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