

NAME: _____ Period: _____

Ch 10 Chemical Quantities Study Guide

/75

Matching

Match each item with the correct statement below.

- a. molar volume
- b. molar mass
- c. atomic mass

- _____ 1. the number of grams of an element that is numerically equal to the atomic mass of the element in amu
- _____ 2. the mass of a mole of any element or compound
- _____ 3. the volume occupied by a mole of any gas at STP

Match each item with the correct statement below.

- a. representative particle
- b. mole
- c. Avogadro's number
- d. percent composition
- e. standard temperature and pressure
- f. empirical formula

- _____ 4. the number of representative particles of a substance present in 1 mole of that substance
- _____ 5. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- _____ 6. the SI unit used to measure amount of substance
- _____ 7. 0°C and 1 atm
- _____ 8. the percent by mass of each element in a compound
- _____ 9. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. What SI unit is used to measure the number of representative particles in a substance?
 - a. kilogram
 - b. ampere
 - c. kelvin
 - d. mole
- _____ 2. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C_3H_7O ?
 - a. $5 \times (6.02 \times 10^{23})$
 - b. 5
 - c. 35
 - d. $35 \times (6.02 \times 10^{23})$
- _____ 3. If a dozen Clementine oranges have a mass of 744 g, what will be the mass of 15 Clementines?
 - a. 11.16 kg
 - b. 595 g
 - c. 896 g
 - d. 930 g
- _____ 4. Which of the following elements exists as a diatomic molecule?
 - a. neon
 - b. lithium
 - c. nitrogen
 - d. sulfur
- _____ 5. Avogadro's number of representative particles is equal to one _____.
 - a. kilogram
 - b. gram
 - c. kelvin
 - d. mole
- _____ 6. Which of the following is equal to Avogadro's number?
 - a. the number of atoms of bromine in 1 mol Br_2
 - b. the number of molecules of chlorine in 0.50 mol Cl_2
 - c. the number of molecules of nitrogen in 1 mol N_2
 - d. the number of atoms in 1 mol CO

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**Barron's Educational Series, Terry L.
Duran**



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