

CHEMICAL EQUILIBRIUM WORKSHEET

On the line at the left, write the letter of the description that best matches each term.

- | | | |
|-------|--------------------------------|--|
| _____ | 1. Equilibrium position | a. used to determine if a reaction has reached equilibrium |
| _____ | 2. Law of chemical equilibrium | b. depends on the initial concentrations of the substances in a reaction |
| _____ | 3. Reaction quotient | c. states that every reaction proceeds to an equilibrium state with a specific K_{eq} |
| _____ | 4. Law of mass action | d. expresses the relative concentration of reactants and products at equilibrium in terms of an equilibrium constant |
| _____ | 5. Equilibrium constant | e. the ratio of product concentration to reactant concentration at equilibrium |

Answer each of the following in the space provided

- What is the equilibrium expression for the equation $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$?
- What is the equilibrium expression for the equation $NH_4Cl(s) \rightleftharpoons NH_3(g) + HCl(g)$?
- What is the equilibrium expression for the equation $As_4O_6(s) + 6C(s) \rightleftharpoons As_4(g) + 6CO(g)$?
- What is the equilibrium expression for the equation $SnO_2(s) + 2CO(g) \rightleftharpoons Sn(s) + 2CO_2(g)$?
- What is the equilibrium expression for the equation $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$?
- For the reaction $2CO(g) \rightleftharpoons C(s) + CO_2(g)$, $K_{eq} = 7.7 \times 10^{-15}$. At a particular time, the following concentrations are measured: $[CO]=0.034$ M, $[CO_2]=3.6 \times 10^{-17}$ M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
- For the reaction $N_2O_4(g) \rightleftharpoons 2NO_2(g)$, $K_{eq} = 0.2$. At a particular time, the following concentrations are measured: $[N_2O_4]=2.0$ M, $[NO_2] = 0.2$ M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
- For the reaction $2ICl(g) \rightleftharpoons I_2(g) + Cl_2(g)$, $K_{eq} = 0.11$. At a particular time, the following concentrations are measured: $[ICl]=2.5$ M, $[I_2] = 2.0$ M, $[Cl_2]= 1.2$ M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
- At 340°C , $K_{eq} = 0.064$ for the reaction $Fe_2O_3(s) + 3H_2(g) \rightleftharpoons 2Fe(s) + 3H_2O(g)$. Given that $[H_2]=0.45$ M and $[H_2O]=0.37$ M, find Q and predict how the reaction will proceed.

Chemical Equilibrium Expression Answer Key

R. Winston Revie



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Chemical Thermodynamics Siddharth Venkatesh, 2025-02-20 Chemical Thermodynamics The Essentials offers a comprehensive and accessible exploration of the fundamental principles and practical applications of thermodynamics in chemical systems Designed for students researchers and professionals this book delves into the energetic underpinnings of chemical reactions and processes Covering basic principles to advanced topics like phase equilibria and chemical kinetics each chapter provides clear explanations illustrative examples and practical applications The book adopts a rigorous

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Principles of Chemical Engineering Processes

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web jun 6 2014 at the top left of the diagram we see the cylinder selector switch most vendors will use one tachometer head to cover a variety of engines make sure this switch is set to the number of cylinders for your engine the right upper l point indicates the internal gauge light

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web jul 11 2010 1 i have a mercruiser 3 0 inboard boat engine i need to know how to wire the ignition switch which color wires go where it has 4 prongs four wires one thick orange one red one thin purple and one yellow any advice the ignition was recently changed but i think they may have wired it wrong thanks achris more fish than

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the positive and negative terminals on the ignition switch disconnect the negative cable from the battery connect the positive terminal of the ignition switch to the positive terminal of the battery

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web sep 13 2013 it's not wired up the engine harness carries the wires from the throttle control to the engine and based on the wiring diagram i found most of those wires are for the key ignition and i guess a portion is for the tach not sure how that works the wires on the internal wiring harness are horrible

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web mar 13 2008 i want to see if someone can give me some info on a wiring diagram for a mercury ignition switch part 8788107a5 the original switch has terminal markings new switch is molded in plastic and i can't tell which terminals the molded wires come from they are color coded but the wires in the control are kinda faded

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web pull the ignition cable through the hole in the dash find the terminal on the switch that's marked b locate the wire in the ignition cable that's red with a purple stripe along its length solder it to the b terminal with silver core solder and a soldering iron solder the purple wire to the a terminal

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