

## Chapter 2 The Nature of Matter • Section 1 Summary

## Describing Matter

### Key Concepts

- What kinds of properties are used to describe matter?
- What are elements, and how do they relate to compounds?
- What are the properties of a mixture?

**Matter** is anything that has mass and takes up space. **Chemistry** is the study of the properties of matter and how matter changes. In chemistry, a **substance** is a single kind of matter that is pure.

Every form of matter has two kinds of properties—**physical properties** and **chemical properties**. A **physical property** is observed without changing a substance into another substance. Examples of physical properties are hardness, texture, color, and ability to dissolve in water. A **chemical property** is the ability of a substance to change into different substances. Some chemical properties are burning and rusting.

All matter is made up of elements. An **element** is a pure substance that cannot be broken down into any other substance. **Elements** are the simplest substances. Each element is identified by its specific physical and chemical properties. An **atom** is the basic particle that makes up an element. Atoms of most elements can combine with other atoms. A **chemical bond** is the force that holds two atoms together. Atoms often combine to form **molecules**, which are larger particles made of two or more atoms held together by chemical bonds.

When elements are chemically combined, they form **compounds** having properties that are different from those of the uncombined elements. A **compound** is a pure substance made of two or more elements chemically combined in a set ratio. A compound may be represented by a **chemical formula**. A chemical formula shows the elements in the compound and the ratio of atoms. For example, the chemical formula for carbon dioxide is  $\text{CO}_2$ . In carbon dioxide, there are always two oxygen atoms to every one carbon atom.

Elements and compounds are pure substances, but most of the materials you see every day are not. Instead, they are mixtures. A **mixture** is made of two or more substances that are together in the same place, but are not chemically combined. Mixtures differ from compounds in two ways. Each substance in a mixture keeps its individual properties. Also, the parts of a mixture are not combined in a set ratio.

A mixture can be heterogeneous or homogeneous. In a **heterogeneous mixture**, you can see the different parts. The substances in a **homogeneous mixture** are so evenly mixed that you cannot see the different parts. A **solution** is an example of a homogeneous mixture. Air is a solution of nitrogen gas, oxygen gas, plus small amounts of other gases. Unlike compounds, mixtures are easily separated into their components. For example, iron filings can be easily removed from salt with a magnet.

# Concept Review Section Describing Matter Answers

**K Payea**



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## **Table of Contents Concept Review Section Describing Matter Answers**

1. Understanding the eBook Concept Review Section Describing Matter Answers
  - The Rise of Digital Reading Concept Review Section Describing Matter Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Section Describing Matter Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Concept Review Section Describing Matter Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review Section Describing Matter Answers
  - Personalized Recommendations
  - Concept Review Section Describing Matter Answers User Reviews and Ratings
  - Concept Review Section Describing Matter Answers and Bestseller Lists

5. Accessing Concept Review Section Describing Matter Answers Free and Paid eBooks
  - Concept Review Section Describing Matter Answers Public Domain eBooks
  - Concept Review Section Describing Matter Answers eBook Subscription Services
  - Concept Review Section Describing Matter Answers Budget-Friendly Options
6. Navigating Concept Review Section Describing Matter Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Concept Review Section Describing Matter Answers Compatibility with Devices
  - Concept Review Section Describing Matter Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Concept Review Section Describing Matter Answers
  - Highlighting and Note-Taking Concept Review Section Describing Matter Answers
  - Interactive Elements Concept Review Section Describing Matter Answers
8. Staying Engaged with Concept Review Section Describing Matter Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Concept Review Section Describing Matter Answers
9. Balancing eBooks and Physical Books Concept Review Section Describing Matter Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Concept Review Section Describing Matter Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Concept Review Section Describing Matter Answers
  - Setting Reading Goals Concept Review Section Describing Matter Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Concept Review Section Describing Matter Answers
  - Fact-Checking eBook Content of Concept Review Section Describing Matter Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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