

Chapter 17 Thermochemistry Study Guide

17.1 – 17.2 Thermochemical Equations

Honors Chemistry

- Make the following conversions:
 - 444 cal to joules = $1.86 \times 10^3 \text{ J}$
 - 1.8 kJ to joules = $1.8 \times 10^3 \text{ J}$
 - 0.45 kJ to calories = $1.1 \times 10^2 \text{ cal}$
- Classify each of these processes as endothermic or exothermic:
 - condensing steam – **exo**
 - burning alcohol – **exo**
 - evaporating alcohol – **endo**
 - baking a potato – **endo**
- The specific heat capacity for silver is $0.24 \text{ J/g}^\circ\text{C}$. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(107.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

- It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

- A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C . The equilibrium temperature in the calorimeter is 21.8°C . Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

- A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C . Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is $0.444 \text{ J/g}^\circ\text{C}$.

$$\begin{array}{rcl} -15.0(.444)(T_f - 100.0) & = & 55.0(4.184)(T_f - 23.0) \\ -6.66T_f + 666 & = & 230.12T_f - 5292.76 \\ 5958.76 & = & 236.78 T_f \\ \mathbf{25.2^\circ\text{C}} & = & T_f \end{array}$$

- Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at 95°C and then adds 15.8 L of water at 75°C , what will the final temperature of the water be?

$$\begin{array}{rcl} -20(T_f - 95) & = & 15.8(T_f - 75) \\ -20T_f + 1900 & = & 15.8 T_f - 1185 \\ 3085 & = & 35.8 T_f \end{array}$$

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