

# Compounds and Their Bonds

9

## A. Ionic Transfer of Electrons

| Element  | Atomic # | 1. Electron Configuration of Atom | 2. Electron-Dot Symbol         | 3. Loss or Gain       | 4. Electron Configuration of Ion | 5. Ionic Charge | 6. Symbol of Ion | 7. Name of Ion |
|----------|----------|-----------------------------------|--------------------------------|-----------------------|----------------------------------|-----------------|------------------|----------------|
| Sodium   | 11       | $1s^2 2s^2 2p^6 3s^1$             | $\cdot \text{Na} \cdot$        | Loss 1 e <sup>-</sup> | $1s^2 2s^2 2p^6$                 | 1 <sup>+</sup>  | $\text{Na}^+$    | Sodium ion     |
| Nitrogen | 7        | $1s^2 2s^2 2p^3$                  | $\cdot \ddot{\text{N}} \cdot$  | Gain 3 e <sup>-</sup> | $1s^2 2s^2 2p^6$                 | 3 <sup>-</sup>  | $\text{N}^{3-}$  | Nitride ion    |
| Aluminum | 13       | $1s^2 2s^2 2p^6 3s^2 3p^1$        | $\cdot \ddot{\text{Al}} \cdot$ | Loss 3 e <sup>-</sup> | $1s^2 2s^2 2p^6$                 | 3 <sup>+</sup>  | $\text{Al}^{3+}$ | Aluminum ion   |
| Chlorine | 17       | $1s^2 2s^2 2p^6 3s^2 3p^5$        | $\cdot \ddot{\text{Cl}} \cdot$ | Gain 1 e <sup>-</sup> | $1s^2 2s^2 2p^6 3s^2 3p^6$       | 1 <sup>-</sup>  | $\text{Cl}^-$    | Chloride ion   |
| Calcium  | 20       | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$   | $\text{Ca} \cdot$              | Loss 2 e <sup>-</sup> | $1s^2 2s^2 2p^6 3s^2 3p^6$       | 2 <sup>+</sup>  | $\text{Ca}^{2+}$ | Calcium ion    |
| Oxygen   | 8        | $1s^2 2s^2 2p^4$                  | $\cdot \ddot{\text{O}} \cdot$  | Gain 2 e <sup>-</sup> | $1s^2 2s^2 2p^6$                 | 2 <sup>-</sup>  | $\text{O}^{2-}$  | Oxide ion      |

## B. Ionic Compounds and Formulas

### 1. Physical properties

| Compound              | Appearance                | Density                | Melting Point |
|-----------------------|---------------------------|------------------------|---------------|
| Sodium chloride, NaCl | White, solid, crystalline | 2.16 g/cm <sup>3</sup> | 801°C         |

### 2. Formulas of ionic compounds

| Name              | Positive Ion     | Negative Ion     | Formula                 |
|-------------------|------------------|------------------|-------------------------|
| Lithium iodide    | $\text{Li}^{+1}$ | $\text{I}^{-1}$  | LiI                     |
| Aluminum oxide    | $\text{Al}^{3+}$ | $\text{O}^{2-}$  | $\text{Al}_2\text{O}_3$ |
| Calcium sulfide   | $\text{Ca}^{2+}$ | $\text{S}^{2-}$  | $\text{CaS}$            |
| Magnesium bromide | $\text{Mg}^{2+}$ | $\text{Br}^{-1}$ | $\text{MgBr}_2$         |
| Potassium nitride | $\text{K}^{+1}$  | $\text{N}^{3-}$  | $\text{K}_3\text{N}$    |
| Sodium fluoride   | $\text{Na}^{+1}$ | $\text{F}^{-1}$  | $\text{NaF}$            |

### 3. Names of ionic compounds

| Formula                 | Positive Ion     | Negative Ion     | Name                |
|-------------------------|------------------|------------------|---------------------|
| $\text{K}_2\text{S}$    | $\text{K}^{+1}$  | $\text{S}^{2-}$  | Potassium sulfide   |
| $\text{BaF}_2$          | $\text{Ba}^{2+}$ | $\text{F}^{-1}$  | Barium fluoride     |
| $\text{MgO}$            | $\text{Mg}^{2+}$ | $\text{O}^{2-}$  | Magnesium oxide     |
| $\text{Na}_3\text{N}$   | $\text{Na}^{+1}$ | $\text{N}^{3-}$  | Sodium nitride      |
| $\text{AlCl}_3$         | $\text{Al}^{3+}$ | $\text{Cl}^{-1}$ | Aluminum chloride   |
| $\text{Mg}_3\text{P}_2$ | $\text{Mg}^{2+}$ | $\text{P}^{3-}$  | Magnesium phosphide |

# Compounds Their Formulas Lab 7 Answers

**H Kauffman**



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