

Common Nonmetallic ROOTS

H = hydr-	B = bor-
C = carb-	N = nitr-
O = ox-	F = fluor-
Si = silic-	P = phosph-
S = sulf-	Cl = chlor-
As = arsen-	Se = selen-
Br = brom-	Sb = antimony-
I = iod-	



Common NONMETAL Ions

S^{2-} sulfide	Br^{-} bromide
F^{-} fluoride	H^{-} hydride
Cl^{-} chloride	I^{-} iodide
P^{3-} phosphide	O^{2-} oxide
N^{3-} nitride	Se^{2-} selenide

Greek PREFIXES

mono- one	hexa- six
di- two	hepta- seven
tri- three	octa- eight
tetra- four	nona- nine
penta- five	deca- ten

Common METAL Ions

Cu^{+} copper (I) (cuprous)	Mn^{2+} manganese (II)
Cu^{2+} copper (II) (cupric)	Ni^{2+} nickel (II)
Fe^{2+} iron (II) (ferrous)	Cr^{2+} chromium (II)
Fe^{3+} iron (III) (ferric)	Cr^{3+} chromium (III)
Hg_2^{2+} mercury (I) (mercurious)	Co^{2+} cobalt (II)
Hg^{2+} mercury (II) (mercuric)	
Pb^{2+} lead (II) (plumbous)	
Pb^{4+} lead (IV) (plumbic)	
Sn^{2+} tin (II) (stannous)	
Sn^{4+} tin (IV) (stannic)	

Polyatomic Ions

BO_3^{3-} borate ion	
BrO_3^{-} bromate ion	
BrO_4^{-} perbromate	
$B_4O_7^{2-}$ tetraborate	
$C_2H_3O_2^{-}$ acetate ion (CH_3COO^{-})	
$C_6H_5O_7^{3-}$ citrate ion	
ClO^{-} hypochlorite ion	
ClO_2^{-} chlorite ion	
ClO_3^{-} chlorate ion	
ClO_4^{-} perchlorate ion	
CN^{-} cyanide ion	
CO_3^{2-} carbonate ion	
$C_2O_4^{2-}$ oxalate ion	
CrO_4^{2-} chromate ion	
$Cr_2O_7^{2-}$ dichromate ion	
HPO_4^{2-} hydrogen phosphate ion	
HSO_4^{-} hydrogen sulfate ion	
HCO_3^{-} hydrogen carbonate ion	
$H_2PO_4^{-}$ dihydrogen phosphate ion	
IO^{-} hypiodite ion	
IO_2^{-} iodite ion	
IO_3^{-} iodate ion	
IO_4^{-} periodate ion	
MnO_4^{-} permanganate ion	
NH_4^{+} ammonium ion	SeO_3^{2-} selenite ion
NO_2^{-} nitrite ion	SeO_4^{2-} selenate ion
NO_3^{-} nitrate ion	SiO_4^{4-} silicate ion
O_2^{2-} peroxide ion	$S_2O_3^{2-}$ thiosulfate ion
OH^{-} hydroxide ion	SO_3^{2-} sulfite ion
PO_3^{3-} phosphite ion	SO_4^{2-} sulfate ion
PO_4^{3-} phosphate ion	SCN^{-} thiocyanate ion

Compound Names Cheat Sheet

Peter J. Mikulecky, Chris Hren



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