



"You can think about stoichiometry chemistry"
 • You can think of molecules as it is as the ingredients to create the product (recipe)
 • Molecules are made from 10 million of a molecule. But for the grams, molecules are made, not for more



5 MORE STOICHIOMETRY SAMPLE PROBLEM

If you have 120 grams glucose and an excess of absolute water, how many grams of water can you make?

pH



Acid Solution
 Contains
 fewer gas
 longer
 battery acid
 gastric acid

Acid solutions have a



Neutral Solution
 Contains
 neither
 cooking oil
 sugar solution

Neutral solutions have a pH



Basic Solution
 Contains
 fewer gas
 longer
 battery acid
 gastric acid

Basic solutions have a

Empirical formula mass $\times n$
 = Molar Mass

Empirical formula CH_2O_2
 multiply subscripts by n
 Molecular formula $\text{C}_2\text{H}_4\text{O}_4$

It has a molar mass of 108.0 grams. Determine the molecular formula. Use the empirical formula from the previous page.

CHEMISTRY UNIT 9: STOICHIOMETRY DOODLE DIAGRAMS



Reactant
 Reactant 1 $\rightarrow$ 1 molecule
 Reactant 2 $\rightarrow$ 2 molecules
 1. Identify which is the limiting reactant.
 Because you only have 2 molecules to react with 10 limiting reactant in this reaction.
 2. Write the product for reactant molecules from the chemical reaction information.
 3. Use the factor label method to calculate the amount of product that is possible to produce using the given reactant amount.

Limiting Reactant

Calculate the volume of the concentrated solution needed that contains that number of moles.

$$\left(\frac{10.0 \text{ moles glucose}}{1} \right) \left(\frac{10.0 \text{ L}}{10.0 \text{ moles glucose}} \right) = 100.0 \text{ L of concentrated solution needed}$$

Now the volume of concentrated solution needed is in graduated cylinder or beaker and add water until the volume comes to the desired volume for this dilute solution.

Formula
 Formula represents the atoms of each type of atom in a compound by molecular and empirical formula. An example is water. Water's atoms are H₂O and water's molecular formula is H₂O.
 Empirical formula can be calculated using that the ratio of each element in a compound composition.

Compound Name
 Glucose

Chemical Formula
 $\text{C}_6\text{H}_{12}\text{O}_6$

Structural Formula
 Structural formula is used to display the individual bonds of each atom in the molecule with the three-dimensional structure that is more similar to the ball-and-stick model.

Chemistry Semester 1 Unit 9 Stoichiometry Notes

Hugh Munby



Chemistry Semester 1 Unit 9 Stoichiometry Notes:

Stanford University Bulletin Stanford University,1998 Teachers And Teaching Hugh Munby,2004-08-02 The intention of this book is to develop an increased awareness of the place of professional practice in the realms of research in teaching The chapters investigate from an international perspective the emerging reflective methods of collaboration between practitioners and researchers appreciation of teachers and teaching and greater understanding of what they aim to promote

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