

When stoichiometric equivalent moles of acid and base react, they are said to be at the *equivalence point*. The process of reaching the equivalence point with acids and bases is called *neutralization*.

In order to know when the equivalence point is reached, an indicator is added to the titration vessel. An *indicator* is a compound that will react with one of the titration species and for which a color change is observed very near the equivalence point, when the appropriate indicator has been selected. The indicator used in this experiment is called phenolphthalein which turns from clear in acid solution to a pink color when excess base is present.

Sample calculation

If 0.558g of potassium hydrogen phthalate (KHP) is dissolved in 50 mL of deionized water in a flask, what is the concentration of a NaOH solution if 12.78 mL are required to titrate the potassium hydrogen phthalate to the equivalence point?

- a) Determine the moles of potassium hydrogen phthalate (KHP) in the flask using mass and molar mass:

$$\text{moles KHP} = \text{mass KHP} \times \frac{1}{M_{\text{KHP}}} = 0.558\text{g} \times \frac{1 \text{ mol}_{\text{KHP}}}{204.23\text{g}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{KHP}}$$

- b) Using the stoichiometric relationship in the balanced chemical reaction (Eq.1), determine the moles of base required to reach the equivalence point:

$$\text{mole NaOH} = \text{mol}_{\text{KHP}} \times \frac{1 \text{ mol}_{\text{NaOH}}}{1 \text{ mol}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{KHP}} \times \frac{1 \text{ mol}_{\text{NaOH}}}{1 \text{ mol}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{NaOH}}$$

- c) Determine the concentration of NaOH by dividing the number of mole by volume of NaOH solution added:

$$\text{molarity of NaOH} = \frac{0.00273 \text{ mol}_{\text{NaOH}}}{0.01278 \text{ L}_{\text{NaOH solution}}} = 0.214 \text{ M NaOH}$$

After the $\approx 0.3 \text{ M}$ NaOH solution has been standardized (its accurate average molarity determined by titration), it can be referred to as a standardized solution and its average molarity from the three titrations should be written on the bottle. It, in turn, will be employed to determine the acetic acid content in household vinegar. Like potassium hydrogen phthalate, acetic acid (ethanoic acid) has one acidic proton, H^+ . Because the acid is monoprotic, it will react one-to-one with NaOH (Eq. 2).



This time you will be starting with a dilute solution of acetic acid in the flask instead of a solid, and you will know the accurate molarity of the NaOH solution (the average of the three values determined). After the titration, knowing the volume of NaOH used in the titration, you can determine the mole of NaOH needed to reach the equivalence point with the acetic acid:

$$\text{moles NaOH} = M_{\text{NaOH}} \left(\frac{\text{mol}}{\text{L}} \right) \times V_{\text{NaOH}} (\text{L})$$

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Table of Contents Chemistry 1211 Lab Manual Uga

1. Understanding the eBook Chemistry 1211 Lab Manual Uga
 - The Rise of Digital Reading Chemistry 1211 Lab Manual Uga
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 1211 Lab Manual Uga
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 1211 Lab Manual Uga
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 1211 Lab Manual Uga
 - Personalized Recommendations
 - Chemistry 1211 Lab Manual Uga User Reviews and Ratings
 - Chemistry 1211 Lab Manual Uga and Bestseller Lists

5. Accessing Chemistry 1211 Lab Manual Uga Free and Paid eBooks
 - Chemistry 1211 Lab Manual Uga Public Domain eBooks
 - Chemistry 1211 Lab Manual Uga eBook Subscription Services
 - Chemistry 1211 Lab Manual Uga Budget-Friendly Options
6. Navigating Chemistry 1211 Lab Manual Uga eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 1211 Lab Manual Uga Compatibility with Devices
 - Chemistry 1211 Lab Manual Uga Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 1211 Lab Manual Uga
 - Highlighting and Note-Taking Chemistry 1211 Lab Manual Uga
 - Interactive Elements Chemistry 1211 Lab Manual Uga
8. Staying Engaged with Chemistry 1211 Lab Manual Uga
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 1211 Lab Manual Uga
9. Balancing eBooks and Physical Books Chemistry 1211 Lab Manual Uga
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 1211 Lab Manual Uga
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 1211 Lab Manual Uga
 - Setting Reading Goals Chemistry 1211 Lab Manual Uga
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 1211 Lab Manual Uga
 - Fact-Checking eBook Content of Chemistry 1211 Lab Manual Uga
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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