Name:	TKE	[Y

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I. Dilution Problems $M_1V_1 = M_2V_2$

 A student wishes to prepare 250. mL of a 0.15 M NaOH solution from a 6.00M NaOH stock solution. What yolume of the stock solution should she use?

How much water must be added to a 2.50 M HCl solution to obtain 500, mL of a 0.100M HCl (2.50M)(x ml) = (0.100 M) (500. ml) solution?

3. What is the final volume of a 0.15 M solution prepared from 25.0 mL of a 6.0 M sodium M=6.0 M M2=0.15 M (6.0 M)(25.0 mL) = (0.15 M) (x ML)

II. Molarity Problems
$$Molarity(M) = \frac{mol}{L}$$

4. What is the molarity of a solution that contains 12.5g NaC2H3O2 (Sodium Acetate) dissolved M = (0.1524 mal) in 0.250 L?

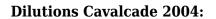
5. What mass of citric acid (molar mass = 192.14 g/mol) would be needed to make 0.10 L of

What mass of sodium hydroxide is needed to make 0.500 L of 2.50 M NaOH?

x = 0.50 mous

Dilutions Cavalcade 2004

Rosina Ehmann



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