

CHEMISTRY



Periodic Table of Elements

1																	18																		
H 1.008																	Ar 39.948																		
2																	36																		
He 4.0026																	Kr 83.80																		
3	4															54	86																		
Li 6.941	Be 9.0122															Xe 131.29	Rn 222																		
5	6															72	84																		
Na 22.990	Mg 24.305															Hf 178.49	Po 209																		
9	10															80	118																		
F 18.998	Ne 20.180															Hg 200.59	Og 294																		
11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36										
K 39.098	Ca 40.078	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72
Rb 85.468	Sr 87.62	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Cs	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	
57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	
Cs 132.905	Ba 137.327	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	Fr	Ra		
87	88	89	90	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120	121	
Fr 223	Ra 226	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo	Uu1	Uu2		
117	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	137	138	139	140	141	142	143	144	145	146	147	148	149	150	151	
Ts 288	Ds 289	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	Fl	Mc	Lv	Nh	

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
 - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatomic:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex:** Central ion + Solvent
 - Transition metal compounds:**
 - Coordination compounds: Non-bonding ligands
 - Organometallics: Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number (Z) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radius:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

- Missing elements were predicted using periodicity
- Columns 1: Alkali metals
- Columns 2: Alkaline earth metals
- First transition series: Titanium-Copper
- Column 17: Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18: Rare gases: helium, neon, argon, krypton, xenon, and radon
- Artificial: $Z = 90$ or above; radioactive
- Lanthanides: f-subshell elements; starts with $Z = 58$



Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell
5 orbitals

d_{xy}

d_{yz}

d_{zx}

$d_{x^2-y^2}$

d_{z^2}

d_{xy}

d_{yz}

d_{zx}

$d_{x^2-y^2}$

d_{z^2}

d_{xy}

d_{yz}

d_{zx}

$d_{x^2-y^2}$

d_{z^2}

d_{xy}

d_{yz}

d_{zx}

$d_{x^2-y^2}$

d_{z^2}

d_{xy}

d_{yz}

d_{zx}

$d_{x^2-y^2}$

d_{z^2}

d_{xy}

d_{yz}

d_{zx}

Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge $-q$, $Z = q$ electrons; for a cation of charge $+q$, $Z = q$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)

First Ionization Potential vs. Z



Atomic Radius vs. Z



Atomic Quantum Numbers & Orbitals

- n : principle**
 - quantum number (orbital direction)
 - For each n , the possible l values are $0, 1, \dots, n-1$
 - For each l , the possible m_l values are $-l, \dots, 0, \dots, +l$
 - m_l has two possible values: $+1/2$ and $-1/2$ (spin up, spin down)
 - Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell
- l : angular momentum (orbital shape)**
 - $l = 0$, 1 type
 - $l = 1$, 3 types
 - $l = 2$, 5 types
 - $l = 3$, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



n	l	m_l	m_s	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	0	$-\frac{1}{2}$	N
2	1	+1	$+\frac{1}{2}$	O
2	1	+1	$-\frac{1}{2}$	F
2	1	+1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up, $m_s = +1/2$, and one with spin down, $m_s = -1/2$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
 - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

Chemistry Section 3 Study Guide

Arshad Iqbal



Chemistry Section 3 Study Guide:

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Table of Contents Chemistry Section 3 Study Guide

1. Understanding the eBook Chemistry Section 3 Study Guide
 - The Rise of Digital Reading Chemistry Section 3 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Section 3 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Section 3 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Section 3 Study Guide
 - Personalized Recommendations
 - Chemistry Section 3 Study Guide User Reviews and Ratings
 - Chemistry Section 3 Study Guide and Bestseller Lists
5. Accessing Chemistry Section 3 Study Guide Free and Paid eBooks
 - Chemistry Section 3 Study Guide Public Domain eBooks
 - Chemistry Section 3 Study Guide eBook Subscription Services
 - Chemistry Section 3 Study Guide Budget-Friendly Options
6. Navigating Chemistry Section 3 Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Section 3 Study Guide Compatibility with Devices
- Chemistry Section 3 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Section 3 Study Guide
 - Highlighting and Note-Taking Chemistry Section 3 Study Guide
 - Interactive Elements Chemistry Section 3 Study Guide
- 8. Staying Engaged with Chemistry Section 3 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Section 3 Study Guide
- 9. Balancing eBooks and Physical Books Chemistry Section 3 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Section 3 Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Section 3 Study Guide
 - Setting Reading Goals Chemistry Section 3 Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Section 3 Study Guide
 - Fact-Checking eBook Content of Chemistry Section 3 Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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