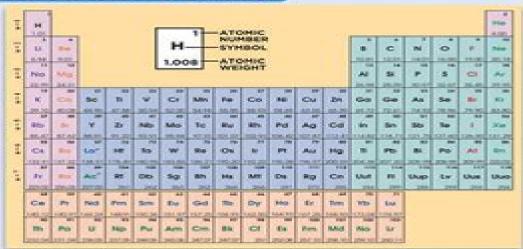




Periodic Table of Elements



as If the property of the property

- Monatomic: Noble gueses: belians, noon, argon, krypton, senon, and radion; monteactive
- Diatomic: Gases: fluorine, chilorine, hydrogen, oxygen, and nitrogen; Liquid: becoming: Solid: inding
- Mutecular solids: Sulfor
- Considered unlider Carbon (diagnosed).
- Metals: Most common type of element; high coordination numbers; worst denne.
- Types of compounds: Lurge variety: due to lonic, covalent, and metallic bonding options.
 - Bomier, Elloctron exchange
 - Covalent: Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex: Central ion # Solvent. Transition metal compounds:
 - Coordination compound: Non-
 - bonding Strands Organometallie: Metal-C bond
 - Complex minerals: Metal exides mental malling attaches
- · Periodicity: The chemical properties of an element depend on valence electrons.
 - Family of elements: Column in. periodic table.
 - The addition of d-orbitals produces. am expranded vertet.

Periodic Trends

- · Chemistry determined by the behavior of valence electrons.
- Atoms or ions got larger as you go down the column; electron shells are added
- Cation and anions (same charec) decrease in size as the atomic number (Z) increases.

- Ionization energies: Noble gases are: the most stable; alkali metals are the easiest to lonize; electron screening lessens Z-effect on orbital energy
- Atomic and looke radii: Not real properties: derive from structural data: divide bond into radii using a consistent and Properties
 - · Ionic rudius: Depends on charge, counter ion, and coordination terratail types).
 - Attomic radius: Based on diatomic gas data or metallic solid structural distant

Development of the Periodic Table

- Missing elements were prodicted using: periodicity
- Culumns In Alkali metals.
- Column 2: Alkaline earth metals
- · First transition series: Titatium-
- Column 17: Welegons: flooring, chiorine, bromine, indine, and autatine
- Culumn 18: Rary penese belium, neon, argon, keypton, senon, and radios
- Actinides: Z = 90 or above; radioactive
- Lanthanides: f-subshell elements: starts with Z - 58



Practical Application













Atomic Structure

- Atomic number (Z): The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge
- -q, Z+q electrons; for a cution of charge +q, Z-q electrons. Mass number (A): A=Z+N, where N = the number of neutrons in the mucleus.
- Instance: Atoms with the same Z. different A: an element can have: one or more instances, any nample contains a number of instances; in practical work, and the average mass number or atomic weight (weighted average of natural isotopes for a given element)





Atomic Quantum Numbers & Orbitals

ms principle.

& angular momentum (setting shape)

max magnetic (orbital direction) max electron spin-■ For each n, the possible / values are 0, 1,..., n - 1

For each L the possible m, values are -L, ..., 0,, ∗I

m, has two possible values: #15 and -15 (spin up, spin down).

· Each value of a denotes a "shell" in the atomic structure; each Edicators a subshelli.

a-carbitals I = 0, 1 type d-orthitals: I = 2, 5 types

p-orbitals I = 1, 3 types fi-orthitals t = 3, 2 types

Many-Electron Atoms

Orbinals and quantum num bers derived for the hydrogen atom are used to describe all munty-electron atoms and lones; a more rigorous treatment determines exact everyslevels and orbital proporties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbara principle:



•	The second second		-	element .	
1	0.0	0.0	100	- 11	I s-cerbital
1	- 0	- 0	-16	3 for	aphorical
2	- 0	0	9.50	1.5	2s-orbital
2	. 0	- 0	-35	Die	aphorical
2	100	1	4.56	- 11	2p-orbitals
2	1	- 0	755	C	
2	1.	+1	996	SN	Bristochial
3	1.0	-1	7486	0	Par Dir Par
2	100	0.7	-36	100	
2:	1	+ 1	-36	C 250mm	

- Electron configuration: Atomic orbital occupancy for an atom.
- Pauli exclusion principle: Each electron has a unique set of quantum numbers; as orbital may hold up to two electrons; one with spin up, m. = +15, and one with spin down, m = -15
- Hand's rule: Electrons fill p. d. and f subshells in a manner to maximize the number of unpoined electrons thalf-filled orbitals); medica but maken head made
- Indication potential (IP): Energy required to remove electron finden are above or lices.
 - · First IP: Removal of first valence electron
- Electromegativity: Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

Chemistry Section 3 Study Guide

Arshad Iqbal

Chemistry Section 3 Study Guide:

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