

Classifying Chemical Reactions

Analyzing and Predicting Products



Introduction

The power of chemical reactions to transform our lives is visible all around us—in our homes, in our cars, even in our bodies. Chemists try to make sense of the great variety of chemical reactions the same way that biologists organize their knowledge of life, by sorting reactions into groups and classifying them. Classifying chemical reactions allows us to predict what chemical reactions will occur when different substances are mixed.

Concepts

- Chemical reactions
- Single vs. double replacement
- Combination vs. decomposition
- Combustion reactions

Background

A chemical reaction is defined as any process in which one or more substances are converted into new substances with different properties. Chemical reactions change the identity of the reacting substance(s) and produce new substances. Observing the properties of the reactants and products is therefore a key step in identifying chemical reactions. Some of the observations that may be associated with a chemical reaction include: (1) release of a gas; (2) formation of a precipitate; (3) color changes; (4) temperature changes; (5) emission or absorption of light. As these observations suggest, chemical reactions can be dynamic and exciting events. The essence of any chemical reaction—reactants being transformed into products—is summarized in the form of a chemical equation. Consider the reaction represented by Equation 1, the burning of natural gas (methane, CH_4) in a laboratory burner.



The reactants—on, more specifically, their formulas—are written on the left side of the equation, the products on the right side of the equation. An arrow represents the direction of the reaction and is read as “yields” or “produces.” Other symbols are used to describe the physical state of the reactants and products and to describe the reaction conditions (see Table 1 on page 2).

Chemical reactions may be classified by considering the number of reactants and products in the reaction, the physical or chemical nature of the reactants and products, and the rearrangement of atoms in the conversion of the reactants into products (see Table 2 on page 2).

Different textbooks may classify reactions differently. Some authors treat all single replacement reactions within a much larger group of oxidation-reduction reactions. Other authors refer to double-replacement reactions as metathesis or double displacement reactions. Finally, a few authors use neither designation and classify double replacement reactions as either precipitation or acid-base reactions.

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Chemical Reactions and Their Equations Ingo Waldemar Dagobert Hackh,1921 **The Ultimate Chemical Equations Handbook, Student Edition** Jane D. Smith,George R. Hague,2011 **Chemical Reactions** Kaufman George Falk,1920
Chemical Reactions in Non-aqueous Solutions Homer O. Lichtenwalter,1913 *Rates of chemical reactions* R. J. Stanley,1971 Encyclopedia of Chemical Reactions. Vol. 5 C. A. Jacobson, *Encyclopedia of Chemical Reactions* C. A. Jacobson,1951 *Chemical Reactions* ,1995 *Rapid Determination of the Order of Chemical Reactions from Time-ratio Tables* Joseph H. Flynn,1960 A table of ration of reaction times corresponding to 15% 5% 20% 10% 95% 85% reaction for orders of chemical reaction from 10 to 10 is presented Use of this table allows the rapid calculation of the order of chemical reactions that are kinetically simple and involves only one subjective step Methods are discussed for interpreting deviations from constant order resulting from errors in the initial times and concentrations from errors in stoichiometry and from the reaction following a more complicated rate expression **Chemical Reactions at High Pressures** K. E. Weale,1973-01-01

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