

AP Chemistry— CHAPTER 16 STUDY GUIDE— Acid-Base Equilibrium

16.1 Acids and Bases: A Brief Review

- Acids taste sour and cause certain dyes to change color.
- Bases taste bitter and feel soapy.
- Arrhenius concept of acids and bases:
 - An acid is a substance that, when dissolved in water, increases the concentration of H^+ ions.
 - Example: HCl is an acid.
 - An Arrhenius base is a substance that, when dissolved in water, increases the concentration of OH^- ions.
 - Example: NaOH is a base.
- This definition is quite narrow in scope as it limits us to aqueous solutions.

16.2 Brønsted-Lowry Acids and Bases

- We can use a broader, more general definition for acids and bases that is based on the fact that acid-base reactions involve proton transfers.

The H^+ Ion in Water

- The $\text{H}^+(\text{aq})$ ion is simply a proton with no surrounding valence electrons.
- In water, clusters of hydrated $\text{H}^+(\text{aq})$ ions form.
- The simplest cluster is $\text{H}_3\text{O}^+(\text{aq})$.
 - We call this a **hydronium ion**.
 - Larger clusters are also possible (such as H_5O_2^+ and H_9O_4^+).
- Generally we use $\text{H}^+(\text{aq})$ and $\text{H}_3\text{O}^+(\text{aq})$ interchangeably.

Proton-Transfer Reactions

- We will focus our attention on $\text{H}^+(\text{aq})$.
- According to the Arrhenius definitions, an acid increases $[\text{H}^+]$ and a base increases $[\text{OH}^-]$.
- Another definition of acids and bases was proposed by Brønsted and Lowry.
- In the Brønsted-Lowry system, a **Brønsted-Lowry acid** is a species that donates H^+ and a **Brønsted-Lowry base** is a species that accepts H^+ .
 - Therefore a Brønsted-Lowry base does not need to contain OH^- .
 - NH_3 is a Brønsted-Lowry base, but not an Arrhenius base.
- Consider $\text{NH}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$:
 - H_2O donates a proton to ammonia.
 - Therefore, water is acting as an acid.
 - NH_3 accepts a proton from water.
 - Therefore, ammonia is acting as a base.
 - An **amphiprotic substance** can behave either as an acid or as a base.
 - Thus, water is an example of an amphiprotic species.

Conjugate Acid-Base Pairs

- Whatever is left of the acid after the proton is donated is called its conjugate base.
- Similarly, a conjugate acid is formed by adding a proton to the base.
- Consider $\text{HX}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{X}^-(\text{aq})$:
 - HX and X^- differ only in the presence or absence of a proton.
 - They are said to be a **conjugate acid-base pair**.
 - X^- is called the **conjugate base**.
 - After HX (acid) loses its proton it is converted into X^- (base).
 - Therefore HX and X^- are a conjugate acid-base pair.
 - After H_2O (base) gains a proton it is converted into H_3O^+ (acid).
 - H_3O^+ is the **conjugate acid**.
 - Therefore, H_2O and H_3O^+ are a conjugate acid-base pair.

Relative Strengths of Acids and Bases

- The stronger an acid is, the weaker its conjugate base will be.
- We can categorize acids and bases according to their behavior in water.
 1. *Strong acids* completely transfer their protons to water.
 - No undissociated molecules remain in solution.
 - Their conjugate bases have negligible tendencies to become protonated.
 - An example is HCl .

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