

9. Evaluate the experiments that led to the conclusion that electrons are negatively charged particles found in all matter.

The deflection toward positively charged plates demonstrated the negatively charged nature of electrons; the fact that changing the type of electrode or the type of gas used in the cathode ray tube did not affect the ray produced led to the conclusion that electrons are present in all matter.

10. Compare the relative charge and mass of each of the subatomic particles.

Particle	Relative charge	Relative mass
Electron	-1	1/1840
Proton	+1	1
Neutron	0	-1

11. **Calculate** What is the difference expressed in kilograms between the mass of a proton and the mass of an electron?

$$1.673 \times 10^{-24} \text{ g} - 9.11 \times 10^{-28} \text{ g} = 1.672 \times 10^{-24} \text{ g} \times 1 \text{ kg}/1000 \text{ g} = 1.672 \times 10^{-27} \text{ kg}$$

Section 4.3 How Atoms Differ

pages 115–121

Practice Problems

pages 116–121

12. How many protons and electrons are in each atom?
- radon, 86 protons and 86 electrons
 - magnesium, 12 protons and 12 electrons
13. An atom of an element contains 66 electrons. Which element is it?
- dysprosium
14. An atom of an element contains 14 protons. Which element is it?
- silicon

15. **Challenge** Do the atoms shown in the figure to the right have the same atomic number?



Yes, 9

16. Determine the number of protons, electrons, and neutrons for isotopes b through f, in the table above. Name each isotope, and write its symbol.

	Protons and Electrons	Neutrons	Isotope	Symbol
b.	20	26	calcium-46	$^{46}_{20}\text{Ca}$
c.	8	9	oxygen-17	$^{17}_8\text{O}$
d.	26	31	iron-57	$^{57}_{26}\text{Fe}$
e.	30	34	zinc-64	$^{64}_{30}\text{Zn}$
f.	80	124	mercury-204	$^{204}_{80}\text{Hg}$

17. **Challenge** An atom has a mass number of 55. Its number of neutrons is the sum of its atomic number and 5. How many protons, neutrons, and electrons does this atom have? What is the identity of this atom?

25 protons, 25 electrons, 30 neutrons, manganese

18. Boron (B) has two naturally occurring isotopes: boron-10 (abundance = 19.8%, mass = 10.013 amu), boron-11 (abundance = 80.2%, mass = 11.009 amu). Calculate the atomic mass of boron.

$$\text{B-10 } 10.013 \text{ amu} \times 0.198 = 1.98 \text{ amu}$$

$$\text{B-11 } 11.009 \text{ amu} \times 0.802 = 8.83 \text{ amu}$$

$$\text{Atomic mass } 1.98 \text{ amu} + 8.83 \text{ amu} = 10.81 \text{ amu}$$

19. **Challenge** Nitrogen has two naturally occurring isotopes, N-14 and N-15. Its atomic mass is 14.007. Which isotope is more abundant? Explain your answer.

N-14 is more abundant because the atomic mass is closer to 14 than 15.

Ch 4 Guided Atomic Structure Answers

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Rani, Reetika Gulati, Sonal
Chauhan, Maukta Gigras**



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Simple Harmonic Motion

Wave Motion

Electrostatics

Capacitance

Current Electricity

Magnetic Effects of Current

Magnetism

EM Induction and AC

electromagnetic Waves

Ray Optics

Wave Optics

Dual Nature of Matter and Radiation

Atoms

Nuclear Physics and Radioactivity

Electronic Devices

Communication Systems

Chemistry

Matter and Laws of Chemical Combinations

Chemical Equations and Stoichiometry

States of Matter

Gaseous and Liquid States

States of Matter

Solid State

Atomic Structure

Radioactivity and Nuclear chemistry

Chemical Bonding and Molecular Structure

Chemical Thermodynamics

Solutions

Chemical Equilibrium

Ionic Equilibrium

Redox Reactions

Electrochemistry

Chemical Kinetics

Adsorption

Colloidal State

Periodic Classification and Periodic Properties

Principles and Process of Metallurgy

Hydrogen

s p d f Block Elements

Coordination Compounds

Environmental Chemistry

Purification of Organic Compounds

Some Basic Principles of Organic Chemistry

Hydrocarbons

Organic Compounds Containing Halogens

Alcohols

Phenols and Ether

Aldehyde

Ketones and Carboxylic Acid

Organic Compounds Containing Nitrogen

Polymers

Biomolecules

Chemistry in Everyday Life

Biology

The Living World

Biological Classification

Plant Kingdom

Animal Kingdom

Morphology of Flowering Plants

Anatomy of Flowering Plants

Structural Organization in Animals

Cell

Biomolecules

Cell Cycle and Cell Division

Transport in Plants

Mineral Nutrition

Photosynthesis

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Table of Contents Ch 4 Guided Atomic Structure Answers

1. Understanding the eBook Ch 4 Guided Atomic Structure Answers
 - The Rise of Digital Reading Ch 4 Guided Atomic Structure Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 4 Guided Atomic Structure Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 4 Guided Atomic Structure Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 4 Guided Atomic Structure Answers
 - Personalized Recommendations
 - Ch 4 Guided Atomic Structure Answers User Reviews and Ratings
 - Ch 4 Guided Atomic Structure Answers and Bestseller Lists
5. Accessing Ch 4 Guided Atomic Structure Answers Free and Paid eBooks

- Ch 4 Guided Atomic Structure Answers Public Domain eBooks
- Ch 4 Guided Atomic Structure Answers eBook Subscription Services
- Ch 4 Guided Atomic Structure Answers Budget-Friendly Options
- 6. Navigating Ch 4 Guided Atomic Structure Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ch 4 Guided Atomic Structure Answers Compatibility with Devices
 - Ch 4 Guided Atomic Structure Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 4 Guided Atomic Structure Answers
 - Highlighting and Note-Taking Ch 4 Guided Atomic Structure Answers
 - Interactive Elements Ch 4 Guided Atomic Structure Answers
- 8. Staying Engaged with Ch 4 Guided Atomic Structure Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 4 Guided Atomic Structure Answers
- 9. Balancing eBooks and Physical Books Ch 4 Guided Atomic Structure Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 4 Guided Atomic Structure Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 4 Guided Atomic Structure Answers
 - Setting Reading Goals Ch 4 Guided Atomic Structure Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 4 Guided Atomic Structure Answers
 - Fact-Checking eBook Content of Ch 4 Guided Atomic Structure Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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