Chemistry

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(Chapter - 3) (Chemical Kinetics) (Exercises)

(Class 12)

Question 3.1:

From the rate expression for the following reactions, determine their order of reaction and the dimensions of the rate constants.

- (i) 3 NO (g) → N₂O (g) Rate = k[NO]²
- (ii) $H_2O_2(aq) + 3I^-(aq) + 2H^* \rightarrow 2H_2O(1) + I_3^- Rate = k[H_2O_2][I^-]$
- (iii) $CH_3CHO(g) \rightarrow CH_4(g) + CO(g)$ Rate = $k[CH_3CHO]^{\frac{2}{3}}$
- (iv) $C_2H_5Cl(g) \rightarrow C_2H_4(g) + HCl(g)$ Rate = $k[C_2H_5Cl]$

Answer 3.1:

(i) Given rate = k[NO]²

Therefore, order of the reaction = 2

Dimension of $k = \frac{Rate}{|NO|^2} \Longrightarrow \frac{\text{mol } L^{-1}s^{-1}}{(\text{mol } L^{-1})^2} = \frac{\text{mol } L^{-1}s^{-1}}{\text{mol}^2L^{-2}} = L \text{ mol}^{-1}s^{-1}$

(ii) Given rate = $k[H_2O_2][I^-]$

Therefore, order of the reaction = 2

Dimension of $k = \frac{\text{Rate}}{[H_2O_2][\Gamma]} \Longrightarrow \frac{\text{mol } L^{-1}s^{-1}}{(\text{mol } L^{-1})(\text{mol } L^{-1})} = L \text{ mol}^{-1}s^{-1}$

(iii) Given rate = $k[CH_3CHO]^{3/2}$

Therefore, order of reaction= 3/2

Dimension of $k = \frac{\text{Rate}}{[\text{CH}_3\text{CHO}]^{3/2}} \Longrightarrow \frac{\text{mol } L^{-1}s^{-1}}{(\text{mol } L^{-1})^{3/2}} = \frac{\text{mol } L^{-1}s^{-1}}{\text{mol}^{3/2}L^{3/2}} = L^{3/2}\text{mol}^{1/2}s^{-1}$

(iv) Given rate = $k[C_2H_5Cl]$

Therefore, order of the reaction = 1

Dimension of $k = \frac{Rate}{[C_2H_3Cl]} \Rightarrow \frac{mol L^{-1}s^{-1}}{mol L^{-1}} = s^{-1}$

Question 3.2:

For the reaction:

 $2A + B \rightarrow A_2B$

The rate= $k[A][B]^2$ with $k=2.0 \times 10^{-6} \text{mol}^{-2} \text{L}^2 \text{s}^{-1}$. Calculate the initial rate of reaction when

 $[A] = 0.1 \text{ mol } L^{-1}$, $[B] = 0.2 \text{mol } L^{-1}$. Calculate the rate of reaction after [A] is reduced to 0.06 mol L^{-1}

Answer 3.2:

The initial rate of reaction is

 $Rate = k[A][B]^2 \implies (2.0 \times 10^{-6} \text{mol}^{-2} \text{L}^2 \text{s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8.0 \times 10^{-9} \text{mol}^{-2} \text{L}^2 \text{s}^{-1}$

When [A]is reduced from 0.1 mol L-1 to 0.06 mol-1,

The concentration of A reacted = (0.1 - 0.06) mol L⁻¹ = 0.04 mol L⁻¹

Therefore, concentration of B reacted = $\frac{1}{2} \times 0.04 \text{ mol L}^{-1}$

Then concentration of B available, $[B] = (0.2 - 0.02) \text{mol L}^{-1} = 0.18 \text{ mol L}^{-1}$

After [A]is reduced to 0.06 mol L-1, the rate of reaction is given by,

 $Rate = k[A][B]^2 \Longrightarrow (2.0 \times 10^{-6} \text{mol}^{-2} \text{L}^2 \text{s}^{-1})(0.06 \text{ mol L}^{-1})(0.18 \text{ mol L}^{-1})^2 = 3.89 \text{ mol L}^{-1} \text{s}^{-1}$

Question 3.3:

The decomposition of NH₃ on platinum is zero order reaction. What are the rates of production of N₂ and H₂ if $k=2.5~\times~10^{-4}~mol^{-1}~L~s^{-1}$?

Answer 3.3:

The decomposition of NH3 on platinum surface is represented by the following equation.

$$2 \text{ NH}_{3(g)} \xrightarrow{\text{Pt}} N_{2(g)} + 3 \text{ H}_{2(g)}$$

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Superoxide Ion Chemistry and Biological Implications Igor B. Afanas'ev,2024-12-11 The chemical properties of superoxide ion its biological role and the role of other oxygen radicals which arise as a result of its transformations are contained in this text In Volume I the principal reactions of superoxide ion including protonation reactions with proton donors nucleophilic reactions with esters alkyl halides and other compounds electron transfer reactions with quinones and metal complexes are described Basic quantitative data including rate constants and yields for the reactions of superoxide ion of all types are given in tables This volume contains the mechanisms of the generation of oxygen radicals in cells and the interaction of superoxide ion with cell components The role of superoxide ion in lipid peroxidation and destruction of proteins and nucleic acids is explained as well as oxygen radicals in the mechanisms of toxic and therapeutic action of drugs especially anticancer antibiotics In addition the action of superoxide ion and other oxygen radicals on plants micro and macroorganisms is discussed along with the role of oxygen radicals in normal metabolic and pathological processes

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